

## **GRAVIMETRIC ANALYSIS OF A CHLORIDE SALT**

*Date submitted: 14/03/2025*

*By: ()*

*Lab section:*

*TA*

**Purpose:** The purpose of this experience is to determine the quantity of the chloride content in an unknown soluble salt.

### **Theory:**

The assignment in this lab was to find the concentration of chloride ion in the given unknown salt sample. Gravimetric analysis is a quantitative analytical technique where the amount of an analyte is determined by measuring the mass of a solid precipitate.

**Precipitation Reaction of Chloride Ions with Silver Ions:** The reaction between chloride ions and silver ions forms a solid precipitate of silver chloride:



So,  $n(\text{AgCl})$  is equal to  $n(\text{Cl})$

### **Solubility of AgCl (s) in water:**

According to the lab manual section "Theory"  $\text{AgCl}(\text{s})$  is very insoluble, which lead to the completion of the reaction.

$$K_{sp} = [\text{Ag}^+][\text{Cl}^-] = 1.6 \times 10^{-10}$$

### **Calculation for the amount of precipitate lost by washing:**

Even though,  $\text{AgCl}$  is insoluble during the washing process some of the precipitate was lost.

Therefore:  $K_{sp} = [\text{Ag}(\text{aq})^+] \cdot [\text{Cl}(\text{aq})^-] = 1.6 \times 10^{-10}$

$$[x] \cdot [x] = 1.6 \times 10^{-10}$$

$$x = 1.3 \times 10^{-5}$$

-  $x$  is equal to  $1.3 \times 10^{-5}$  moles per litre

### **Precipitation is performed in an acidic environment:**

# Gravimetric Analysis Of Chloride Salt Lab Report

**Yi-Tong Ma**



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## **Table of Contents Gravimetric Analysis Of Chloride Salt Lab Report**

1. Understanding the eBook Gravimetric Analysis Of Chloride Salt Lab Report
  - The Rise of Digital Reading Gravimetric Analysis Of Chloride Salt Lab Report
  - Advantages of eBooks Over Traditional Books
2. Identifying Gravimetric Analysis Of Chloride Salt Lab Report
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Gravimetric Analysis Of Chloride Salt Lab Report
  - User-Friendly Interface
4. Exploring eBook Recommendations from Gravimetric Analysis Of Chloride Salt Lab Report

- Personalized Recommendations
  - Gravimetric Analysis Of Chloride Salt Lab Report User Reviews and Ratings
  - Gravimetric Analysis Of Chloride Salt Lab Report and Bestseller Lists
5. Accessing Gravimetric Analysis Of Chloride Salt Lab Report Free and Paid eBooks
    - Gravimetric Analysis Of Chloride Salt Lab Report Public Domain eBooks
    - Gravimetric Analysis Of Chloride Salt Lab Report eBook Subscription Services
    - Gravimetric Analysis Of Chloride Salt Lab Report Budget-Friendly Options
  6. Navigating Gravimetric Analysis Of Chloride Salt Lab Report eBook Formats
    - ePub, PDF, MOBI, and More
    - Gravimetric Analysis Of Chloride Salt Lab Report Compatibility with Devices
    - Gravimetric Analysis Of Chloride Salt Lab Report Enhanced eBook Features
  7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Gravimetric Analysis Of Chloride Salt Lab Report
    - Highlighting and Note-Taking Gravimetric Analysis Of Chloride Salt Lab Report
    - Interactive Elements Gravimetric Analysis Of Chloride Salt Lab Report
  8. Staying Engaged with Gravimetric Analysis Of Chloride Salt Lab Report
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Gravimetric Analysis Of Chloride Salt Lab Report
  9. Balancing eBooks and Physical Books Gravimetric Analysis Of Chloride Salt Lab Report
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Gravimetric Analysis Of Chloride Salt Lab Report
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Gravimetric Analysis Of Chloride Salt Lab Report
    - Setting Reading Goals Gravimetric Analysis Of Chloride Salt Lab Report
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Gravimetric Analysis Of Chloride Salt Lab Report

- Fact-Checking eBook Content of Gravimetric Analysis Of Chloride Salt Lab Report
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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