



- 1) What is happening to the average kinetic energy of the molecules in the sample during section 2? **remains constant**
- 2) As a substance goes through section (2), what happens to the distance between the particles? **increases**
- 3) What is the name of the process happening during section (4)? **boiling (vaporization)**
- 4) What would be the name of the process happening during section (4) if time were going the other way? **condensation**
- 5) What is the melting point of this substance? **0 °C**
- 6) At what temperature would this sample finish boiling? **100 °C**
- 7) When this substance is melting, the temperature of the ice-water mixture remains constant because:
d. **Heat energy is being converted to kinetic energy**
- 8) When a given quantity of water is heated at a constant rate, the phase change from liquid to gas takes longer than the phase change from solid to liquid because
a. **The heat of vaporization is greater than the heat of fusion**
- 9) The temperature at which a substance in the liquid state freezes is the same as the temperature at which the substance
a. **Melts**
- 10) Is this curve showing an **endothermic** or an **exothermic** process?
- 11) 34°C is equal to **307 K**.
- 12) -128°C is equal to **145 K**.

Heating Curve Problems Key

Siddappa N.Byrareddy



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