

Heat Capacity and Specific Heat Calculations - Practice

A 1.000 g sample of decane, $C_{10}H_{22}$, is ignited inside a bomb calorimeter. The temperature rises from 20.0°C to 78.8°C . Given that the heat capacity of the calorimeter is $810.1 \text{ J}/^\circ\text{C}$, determine the heat (q) for the combustion of decane in units of kJ/mol of decane.

$$(1)(810.1)(58.8)$$

$$q = 47633.9 \text{ J}$$

Calculate the final temperature after 1485 J of energy is added to 16.7 g of water at 23.4°C .

$$q = mC\Delta T$$

$$1485 = (16.7)(4.184)(T_f - 23.4)$$

$$44.65^\circ\text{C}$$

$$1485 = (69.8728)(T_f) - 1635.024$$

$$3120.024 = 69.8728 T_f$$

How much heat is absorbed by a copper penny with a mass of 3.10 g when its temperature rises from -8.0°C to 37°C ?

$$T_f = 37$$

$$T_i = -8.0$$

$$37 - (-8) = 45^\circ\text{C}$$

$$(3.10 \text{ g})(45)(0.385) = 53.71 \text{ J}$$

An unknown metal has a mass of 19.7 g . When $8.20 \times 10^2 \text{ J}$ of heat are added to the sample, the sample temperature changes by 55.5°C . Calculate the specific heat capacity of the unknown metal.

$$mC\Delta T$$

$$m = 19.7$$

$$C =$$

$$\Delta T = 55.5$$

$$q = 820 \text{ J}$$

$$(19.7)(C)(55.5) = 820$$

$$C = 0.75 \text{ J}/\text{g}\cdot^\circ\text{C}$$

Heat Calculations Practice Answers

S Nieto



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