

JEE Main 2020 Paper



Sol. 2



P – OH bonds = 4
 P = O bonds = 2
 P – O – P linkage = 1
 Ans. 4, 2, 1

21. The mole fraction of glucose ($\text{C}_6\text{H}_{12}\text{O}_6$) in an aqueous binary solution is 0.1. The mass percentage of water in it, to the nearest integer, is _____.

Sol. 47 %

$$x_{\text{Glucose}} = 0.1$$

$$\begin{aligned}
 \text{mass\% of glucose} &= \frac{0.1 \times 180}{0.1 \times 180 + 0.9 \times 18} \times 100 \\
 &= \frac{1800}{18 + 16.2} \\
 &= \frac{1800}{34.2} \\
 &= 52.63\%
 \end{aligned}$$

$$\therefore \text{mass \% of H}_2\text{O} = 47\%$$

22. The volume strength of 8.9 M H_2O_2 solution calculated at 273 K and 1 atm is _____. ($R = 0.0821 \text{ L atm K}^{-1} \text{ mol}^{-1}$) (rounded off to the nearest integer)

Sol. 100

$$\begin{aligned}
 \text{Vol. strength} &= \frac{8.9}{2} \times \frac{0.0821 \times 273}{1} \\
 &= 99.73 \\
 &= 100
 \end{aligned}$$

23. An element with molar mass $2.7 \times 10^{-2} \text{ kg mol}^{-1}$ forms a cubic unit cell with edge length 405 pm. If its density is $2.7 \times 10^3 \text{ kg m}^{-3}$, the radius of the element is approximately _____ $\times 10^{-12} \text{ m}$ (to the nearest integer).

Sol. 143

$$\text{Density} = \frac{Z \times \text{GMM}}{N_A \times a^3}$$

$$\begin{aligned}
 2.7 \times 10^3 &= \frac{Z \times 2.7 \times 10^{-2}}{6.023 \times 10^{23} \times (405 \times 10^{-12})^3} \\
 Z &= \frac{6.023 \times 405 \times 405 \times 405 \times 10^{23-26+3+2}}{2.7 \times 10^{-2}} \\
 Z &= 4
 \end{aligned}$$

FCC

$$4R = \sqrt{2} \times a$$

$$R = \frac{405}{2\sqrt{2}} \times 10^{-12} = 143.21 \times 10^{-12} \text{ m}$$

$$= 143 \text{ ans}$$

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Mark Steyvers



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