

Experiment: Solubility Rules and Net Ionic Equations

This is to be written up in its entirety in your laboratory notebook

OBJECTIVES

In this experiment you will mix small quantities of aqueous ionic solutions together and in some cases, observe the formation of precipitates.

OVERVIEW

1. Mix small quantities of different ionic solutions together.
2. Observe the absence or presence of a precipitate and its physical characteristics such as color, amount, etc.
3. Formulate general solubility rules based on experimental observations.
4. Write net ionic equations to describe the precipitation reactions.

SAFETY

- Wear safety goggles at all times.
- Wear shoes that cover the entire foot.
- Never touch, taste or smell chemicals.
- Keep your hands away from your face, eyes, nose and mouth.
- Always wash your hands before leaving the laboratory.

PRELAB QUESTIONS

References: Ebbing, "General Chemistry", 5th ed. pp. 91-94 (Molecular-Ionic Equations), pp. 95-98 (Precip. Reactions)

1. Write the balanced equation for dissolving the following solid ionic salts in water:



2. Write the complete ionic and the balanced net ionic equations for the following reactions:



3. Use the general solubility rules to predict which of the following compounds are soluble in water and which are not.

a. NiCl_2

b. NaOH

c. $\text{Ni}(\text{OH})_2$

d. NaCl

e. $(\text{NH}_4)_2\text{S}$

f. $\text{Pb}(\text{NO}_3)_2$

g. PbS

h. NH_4NO_3

i. CaSO_2

j. K_2CO_3

k. K_2SO_4

l. CaCO_3

m. KBr

n. FeCl_3

o. FeBr_3

p. KCl

Experiment 26 Solubility Rules Answer Key

Jessica J Manson



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Laboratory Manual for Fundamentals of General, Organic, and Biological Chemistry John R. Holum, Ruth C. Denison, 1978

The Magic of Science: a Manual of Amusing and Instructive Scientific Experiments ... James Wylde, 1873 **A Laboratory Manual of Elementary Physical Chemistry** Edward Mack, Wesley George France, 1928 *The magic of science, a manual of easy scientific experiments* James Wylde, 1873 *E3 Chemistry Guided Study Book - 2018 Home Edition (Answer Key Included)* Effiong Eyo, 2017-12-08

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OECD Guidelines for the Testing of Chemicals, Section 1 Test No. 106: Adsorption -- Desorption Using a Batch Equilibrium Method OECD,2000-01-21 This Test Guideline is aimed at estimating the adsorption desorption behaviour of a chemical on different soil types The goal is to obtain a sorption value which can be used to predict partitioning under a variety of environmental conditions to *Addison-Wesley Small-scale Chemistry* Dennis D. Staley,Edward L. Waterman,1995

Journal of the American Chemical Society American Chemical Society,1920 Proceedings of the Society are included in v 1 59 1879 1937 **Chemistry in the Laboratory** James M. Postma,Julian L. Robert,J. Leland Hollenberg,2004-03-12 This clearly written class tested manual has long given students hands on experience covering all the essential topics in general chemistry Stand alone experiments provide all the background introduction necessary to work with any general chemistry text This revised edition offers new experiments and expanded information on applications to real world situations

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