

Moles and Chemical Formulas

11

A. Finding the Simplest Formula

1. Mass of empty crucible + cover
2. Initial appearance of the magnesium
3. Mass of crucible + cover + magnesium
4. Mass of crucible + cover + oxide product

	<u>22.942</u>	g
	<u>23.483</u>	
	<u>23.073</u>	g
	<u>29.317</u>	g

Calculations

5. Mass of magnesium
6. Mass of magnesium compound
7. Mass of oxygen in the product
8. Moles of Mg

	<u>0.2305</u>	g
	<u>6.445</u>	g
	<u>6.245</u>	g
	<u>.0095</u>	g

(Show calculations.)

$$2305 \text{ g/mg} \times \frac{1 \text{ mole mg}}{24.31 \text{ g/mg}} = .009481$$

9. Moles of O

(Show calculations.)

$$6.245 \text{ g} \times \frac{1 \text{ mole}}{16.00} = .3903$$

10. Which number of moles (Mg or O) is smaller

Mole of Mg
moles of Mg (rounded)

Experiment 5 Chemical Formulas Answers

Léo Pomar



Experiment 5 Chemical Formulas Answers:

Unveiling the Magic of Words: A Review of "**Experiment 5 Chemical Formulas Answers**"

In a world defined by information and interconnectivity, the enchanting power of words has acquired unparalleled significance. Their power to kindle emotions, provoke contemplation, and ignite transformative change is actually awe-inspiring. Enter the realm of "**Experiment 5 Chemical Formulas Answers**," a mesmerizing literary masterpiece penned by way of a distinguished author, guiding readers on a profound journey to unravel the secrets and potential hidden within every word. In this critique, we shall delve to the book is central themes, examine its distinctive writing style, and assess its profound effect on the souls of its readers.

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