

## Like Dissolves Like Practice Problems

Using your knowledge of the "like dissolves like" principle of polarity, determine which component will dissolve best in each of the following solvents. Explain your choice for each solvent and the determination.

1) Solvent: methanol (CH<sub>3</sub>OH)

Components: carbon tetrachloride and oxygen dichloride

Both oxygen dichloride and methanol are polar, while carbon tetrachloride is not.

2) Solvent: carbon dioxide

Components: methanol (CH<sub>3</sub>OH) and carbon dioxide

While both components are polar, methanol is considerably more polar than carbon dioxide (both components always have that charge separation, while carbon dioxide, as neutral does not), so not.

3) Solvent: water

Components: carbon tetrachloride and ammonia

Both water and ammonia are polar. Carbon tetrachloride is not.

4) Solvent: hexane (C<sub>6</sub>H<sub>14</sub>)

Components: water and methanol

Both water and methanol are polar. Methanol is not.

# Like Dissolves Like Practice Answer Key

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