

Procedure for making NaOH-Acid (for every lab group of 3-4 people):

- Using your calculations from question #4 above, measure out the correct amount of solid NaOH powder (grams) into the assigned cup to make a 0.1 M solution. **Be careful not to include the mass of the cup in your measurement!**
- Transfer the powder into the assigned "mixing cup".
- Add water into the "mixing cup" until you have 100 mL of solution. Stir with a spoon or stick.
- Observe and taste the solutions you have made. You can have one "designated taster" or you can pour a little into separate little cups for each group member to taste. Record in data table.
- Repeat steps 1-3 in order to make the 0.4 M and 0.7 M solutions.
- Clean-Up. Throw away cups that you drank out of, but save the mixing & measuring cups.

Observations/ Descriptions	0.1 M solution	0.4 M solution	0.7 M solution
Observations (color, smell, taste)	Answers will vary	Answers will vary	Answers will vary

Calculations/ Analysis:

- Which concentration that you tasted was closest to the ideal concentration of NaOH-Acid? What was wrong with each of the other solutions that you made?

Answers will vary. Some solutions will not have enough NaOH-Acid for your taste, while others will have too much.

- What was the solute used in this lab? What was the solvent?

Solute = NaOH-Acid powder Solvent = water

- What is molarity?

Molarity is a measure of the number of moles of solute per liter of solution.

Questions for Discussion: Do your best show the formulae of your work, and **highlight** your answer!

- Write the formulae for the 4 different ways to calculate concentrations:

- Molarity

$$M = \frac{\text{moles}}{\text{volume}}$$

- grams/Liter

$$\frac{\text{grams of solute}}{\text{volume of solution}}$$

- % composition

$$\frac{\text{grams of solute}}{\text{grams of solution}} \times 100$$

- Parts per million

$$\frac{\text{grams of solute}}{\text{grams of solution}} \times 10^6$$

Determine the following concentrations, using the 4 different methods:

(Remember: 1 gram NaOH = 1 mL of 1% = 1000 mg/L)

- If 1.00 moles of NaOH was dissolved in enough water to make 1.00 L of solution:

- Molarity?

$$M = \frac{1.00 \text{ mol}}{1.00 \text{ L}}$$

$$M = 1.000 \text{ M NaOH}$$

- grams/Liter?

$$\frac{40.0 \text{ g NaOH}}{1.00 \text{ L}}$$

$$= 40.0 \text{ g/L}$$

- % composition?

$$\frac{40.0 \text{ g NaOH}}{1000 \text{ g}} \times 100$$

$$= 4.00\% \text{ NaOH}$$

- Parts per million?

$$\frac{40.0 \text{ g NaOH}}{1000 \text{ g}} \times 10^6$$

$$= 40,000 \text{ ppm}$$

- If 20 grams of NaOH was dissolved in enough water to make a 250 mL NaOH solution:

- Molarity?

$$M = \frac{0.50 \text{ mol}}{0.250 \text{ L}}$$

$$M = 2.0 \text{ M NaOH}$$

- grams/Liter?

$$\frac{20 \text{ g NaOH}}{0.250 \text{ L}}$$

$$= 80 \text{ g/L}$$

- % composition?

$$\frac{20 \text{ g NaOH}}{250 \text{ g}} \times 100$$

$$= 8\% \text{ NaOH}$$

- Parts per million?

$$\frac{20 \text{ g NaOH}}{250 \text{ g}} \times 10^6$$

$$= 80,000 \text{ ppm}$$

Kool Aid Lab Answer Key

RS Peters



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