

MINI PROJECTOR

User's Manual

English

Deutsch



◆ The contents of this manual are subject to change without prior notice

Warning

Please read this manual carefully before using projector.

If necessary, please note the Model no. and Series No. of this projector.

- ◆ Don't look straight into the light from the projector lens.
- ◆ Direct burst of intensive light can harm your eyes. Can cause severe damage to eyes of infants.
- ◆ Don't put projector in a damp, humid environment to avoid electrocution.
- ◆ The power on the unit is rated at 90 ~ 240V AC. Be aware of outlets being used.
- ◆ Don't open up or attempt repair projector without necessary support / guide due to higher voltage inside unit. Repair should be performed by a professional technician according to the manual instructions or distributor's suggestion.
- ◆ Please ensure the projector is a safe distance from other heat emitting devices, potentially the service life of projector can be shortened.
- ◆ We strongly suggest you power off projector when use it for hours constantly one time.

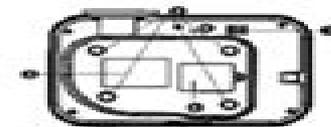
Visual Instructions

◆ Top View



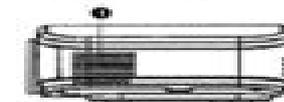
- ① Indicator Light
- ② Power Button

◆ Bottom View



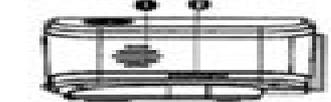
- ① Ventilation Window
- ② Support Base
- ③ Product Code Area
- ④ Dust Cleaner Cover
- ⑤ Adjustable support base

◆ Side View -Left



- ① Ventilation window

◆ Side View -Right



- ① Speaker
- ② Ventilation window

Mini One User Manual

S Marginson



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web donor whereas a base generically represented by b is a proton acceptor when acids and bases combine they swap a proton h_a b h_b strong acids e.g. sulfuric acid have a very pronounced tendency to transfer a proton to a base whereas weak acids

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web understand that parts of an acid base titration be able to determine the K_a or K_b from pH data associated with the titration of a weak acid or base be able to determine the molar mass of a solid monoprotic acid from titration data be able to calculate K_{a1} and K_{a2} for a polyprotic acid by the end of this lab students should be able to

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web answer based on the data the strong acid strong base combination of hydrochloric acid and sodium hydroxide had the neutral ph at the equivalence point being 7 this result is expected because a combination of strong acids and strong bases should result in a neutral at the equivalence point

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web 1 approximately 3 g of hydrated oxalic acid $\text{H}_2\text{C}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$ was weighted to the nearest 0.01 g in a 50 ml beaker 2 about 30 ml of distilled water was added to the beaker to dissolve the oxalic acid 3 the solution was transferred into 250 ml volumetric flask the beaker was rinsed and its contents were poured into the flask

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