

## Chapter 4 Atomic Structure

**Section 4.2 The Structure of an Atom****(pages 108–112)**

*This section compares the properties of three subatomic particles. It also discusses atomic numbers, mass numbers, and isotopes.*

**Reading Strategy (page 108)**

**Monitoring Your Understanding** Before you read, list in the table shown what you know about atoms and what you would like to learn. After you read, list what you have learned. For more information on this Reading Strategy, see the **Reading and Study Skills** in the **Skills and Reference Handbook** at the end of your textbook.

What I Know About Atoms	What I Would Like to Learn	What I Have Learned
Most students will know that atoms are the “building blocks” of matter, and some may know that atoms contain subatomic particles.	Based on the title of the section, students may say that they want to learn more about the structure of atoms.	

**Properties of Subatomic Particles (pages 108–109)**

- What are three subatomic particles?  
a. Protons      b. Electrons      c. Neutrons
- Circle the letter that identifies a subatomic particle with a positive charge.  
a. nucleus      **(b)** proton  
c. neutron      d. electron
- Why did Chadwick conclude that the particles produced by his experiment were neutral in charge? A charged object did not deflect the paths of the particles.  
\_\_\_\_\_  
\_\_\_\_\_

**Comparing Subatomic Particles (pages 109–110)**

- Circle the letters of properties that vary among subatomic particles.  
a. color      **(b)** mass  
**(c)** charge      **(d)** location in the atom
- Circle the letter of the expression that accurately compares the masses of neutrons and protons.  
**(a)** mass of 1 neutron = mass of 1 proton  
b. mass of 2000 neutrons = mass of 1 proton  
c. mass of 1 electron = mass of 1 proton  
d. mass of 1 neutron = mass of 1 electron

# Introduction To Atoms Answer Key

**B Lingard**



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### **Table of Contents Introduction To Atoms Answer Key**

1. Understanding the eBook Introduction To Atoms Answer Key
  - The Rise of Digital Reading Introduction To Atoms Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Introduction To Atoms Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Introduction To Atoms Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Introduction To Atoms Answer Key
  - Personalized Recommendations
  - Introduction To Atoms Answer Key User Reviews and Ratings
  - Introduction To Atoms Answer Key and Bestseller Lists

5. Accessing Introduction To Atoms Answer Key Free and Paid eBooks
  - Introduction To Atoms Answer Key Public Domain eBooks
  - Introduction To Atoms Answer Key eBook Subscription Services
  - Introduction To Atoms Answer Key Budget-Friendly Options
6. Navigating Introduction To Atoms Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Introduction To Atoms Answer Key Compatibility with Devices
  - Introduction To Atoms Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Introduction To Atoms Answer Key
  - Highlighting and Note-Taking Introduction To Atoms Answer Key
  - Interactive Elements Introduction To Atoms Answer Key
8. Staying Engaged with Introduction To Atoms Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Introduction To Atoms Answer Key
9. Balancing eBooks and Physical Books Introduction To Atoms Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Introduction To Atoms Answer Key
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Introduction To Atoms Answer Key
  - Setting Reading Goals Introduction To Atoms Answer Key
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Introduction To Atoms Answer Key
  - Fact-Checking eBook Content of Introduction To Atoms Answer Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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