

# Reaction Kinetics Lab Reporting Worksheet

In science, reporting what has been done in a laboratory setting is incredibly important for communicating, replicating, and validating findings. However, writing scientific reports can be a little overwhelming. There are a set of agreed upon components that the scientific community requires when reporting scientific research. Answer the following questions to describe what occurred during the lab you conducted in Labster. Be sure to use complete sentences and descriptions that fully represent what you experienced. Writing a lab report is less about being correct or incorrect, than it is accurately reporting what happened and why. So, do not worry about reporting data that might seem counterintuitive or unexpected. Focus on clearly communicating what you did and what you observed.

**Write your answers on a new line.**

- What was the title of the lab you completed?
- [Reaction Kinetics: The Essentials](#)

## 1. Topic

- What was the subject you were trying to understand better in the lab?
- Describe the main factors that influence the rate of a chemical reaction
- Give examples of their effects
- Explain the relation between temperature and reaction rate
- Describe the relation between collision theory and activation energy
- Interpret reaction energy diagrams and relate them to energy changes during a reaction
- Suggest a suitable experimental setup for measuring the kinetics of a reaction

## 2. Background Information

- What information from the textbook and classroom is relevant for the subject you were trying to gain a better understanding of in the lab? Identify the concepts and explain how they are related to the lab topic.
- [Temperature](#) - a varying aspect that can change the outcome of a reaction
- [Solvent](#) - a substance in which other materials dissolve to form a solution
- [Catalyst](#) - a substance that enables a chemical reaction to proceed at a faster than usual rate under specific altered conditions than otherwise possible. The substance does not consume itself with this process.
- [Collision theory](#) - effect of concentration and temperature of the on the reaction rate
  - Reactant molecules only react when they collide with the correct orientation and enough energy to overcome the activation energy.
  - Increased collisions therefore increase the reaction rates
- [Activation energy](#) - the minimum energy required to initiate a reaction
  - The lower the activation energy the higher the reaction rate (less to overcome)
- [Exothermic reactions](#) release heat - creating a negative change in H
- [Endothermic reactions](#) absorb heat - creating a positive change in H

# Growth Kinetics Lab Report

**Rachel S Tattersall**



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