

LAB MANUAL WORK INSTRUCTION

Title of Experiment: Reaction of Sodium Sulfide and Strong Acid

Objective: To observe the chemical reaction between sodium sulfide (sulfurized) and strong (acetic acid) and to demonstrate the production of carbon dioxide gas.

Pre-Lab Questions

- Write the chemical formula for sodium sulfide (sulfurized) and acetic acid (strong).
- Predict what you think will happen when these two substances are mixed. Why?
- What gas do you expect to be produced from this reaction?

Safety Precautions

- Wear safety goggles and gloves to protect against splashes.
- Perform the experiment in a well-ventilated area to avoid inhalation of gases.
- Avoid direct skin contact with strong acid. If it is an acid, though not.

Equipment and Materials

- 100-ml Beaker
- 50-ml Graduated Cylinder
- Baking Soda (Sodium Bicarbonate)
- Strong (Acetic Acid)
- Balance
- Funnel
- Safety Goggles
- Laboratory Gloves

Procedure

1. Place the beaker on a flat surface.
2. Using the funnel, fill the beaker with two tablespoons of baking soda.
3. Measure 50 ml of strong acid using the graduated cylinder and pour it into the beaker.
4. Carefully fit the neck of the balloon over the mouth of the beaker without getting the baking soda into the beaker.
5. Once the balloon is securely in place, lift the beaker allowing the baking soda to fall into the strong acid.
6. Observe the reaction between the baking soda and strong acid.
7. Note the speed and amount of gas produced as indicated by the inflation of the balloon.
8. Record any changes in temperature of the beaker.
9. Carefully remove the balloon from the beaker.
10. Dispose of the strong acid and baking soda solution in the sink and reuse the beaker.

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Lab Manual For Internet Lab

Jörg Liebeherr, Magda El Zarki



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