

Ionic Bonding Worksheet

For each pair of elements below draw an atomic diagram showing electrons in different energy levels. Draw arrows to show where the outer electrons will go during a chemical reaction, then draw the resulting compound. Finally, fill in the table below each reaction. Refer to the sample shown.



Atoms	Valence electrons	Electron transfer from/to each atom	Ions formed in the product
Li	1ve	transfers 1ve ⁻ to F	Li^+
F	7ve	accepts 1ve ⁻ from Li	F^-

Reactions	Atoms	Valence electrons	Electron transfer from/to each atom	Ions formed in the product
1) $\text{Li} + \text{Cl} \Rightarrow \text{LiCl}$ 	Li	1ve	donate $1e^-$	Li^+
	Cl	1ve	accepts $1e^-$	Cl^-
2) $\text{Ca} + \text{O} \Rightarrow \text{CaO}$ 	Ca	2ve	donates $2e^-$	Ca^{+2}
	O	6ve	accepts $2e^-$	O^{-2}
3) $\text{Be} + \text{F} \Rightarrow \text{BeF}_2$ 	Be	2ve	donates $2e^-$ to each F	Be^{+2}
	F	7ve	each F accepts 1ve	F^-
4) $\text{Mg} + \text{S} \Rightarrow \text{MgS}$ 	Mg	2ve	donates $2e^-$ to S	Mg^{+2}
	S	6ve	accepts $2e^-$	S^{-2}
5) $\text{K} + \text{F} \Rightarrow \text{KF}$ 	K	1ve	donates 1ve ⁻	K^+
	F	7ve	accepts 1ve ⁻	F^-

Ionic Bonds And Notetaking Guide Answer

Clemens Wendtner



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