

Boyle's Law	Charles' Law	Guy-Lassac's Law	Combined Gas Law
For a given mass of gas at constant temperature, the volume of a gas varies inversely with pressure	The volume of a fixed mass of gas is directly proportional to its Kelvin temperature if the pressure is kept constant.	The pressure of a gas is directly proportional to the Kelvin temperature if the volume is kept constant.	Combines Boyle's, Charles', and the Temperature-Pressure relationship into one equation. Each of these laws can be derived from this law.
$PV = k$ $P_1V_1 = P_2V_2$	$\frac{V}{T} = k$ $V_1T_2 = V_2T_1$ $\frac{V_1}{T_1} = \frac{V_2}{T_2}$	$\frac{P}{T} = k$ $P_1T_2 = P_2T_1$ $\frac{P_1}{T_1} = \frac{P_2}{T_2}$	$\frac{PV}{T} = k$ $V_1P_1T_2 = V_2P_2T_1$ $\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2}$

Dalton's Law	Ideal Gas Law	Graham's Law
At constant volume and temperature, the total pressure exerted by a mixture of gases is equal to the sum of the pressures exerted by each gas.	The Ideal Gas Law relates the pressure, temperature, volume, and mass of a gas through the gas constant "R".	The rate of effusion/diffusion of two gases (A and B) are inversely proportional to the square roots of their formula masses. [It can be a ratio of molecular speeds, effusion /diffusion times, distance traveled by molecules, or amount of gas effused]
$P_{total} = P_1 + P_2 + P_3 + \dots P_n$	$PV = nRT$	$\frac{\text{Rate}_A}{\text{Rate}_B} = \sqrt{\frac{\text{molar mass}_B}{\text{molar mass}_A}}$

Abbreviations	Standard Conditions
atm = atmosphere mm Hg = millimeters of mercury torr = another name for mm Hg Pa = Pascal kPa = kilopascal K = Kelvin °C = degrees Celsius	0°C = 273 K 1.00 atm = 760.0 mm Hg = 76 cm Hg = 101.325 kPa = 101, 325 Pa = 29.9 in Hg
Conversions	Gas Law's Equation Symbols
$K = ^\circ C + 273$ $F = 1.8C + 32$ $C = \frac{F - 32}{1.8}$ 1 cm ³ (cubic centimeter) = 1 mL (milliliter) 1 dm ³ (cubic decimeter) = 1 L (liter) = 1000 mL	Subscript (1) = old condition or initial condition Subscript (2) = new condition or final condition Temperature must be in Kelvins n = number of moles = grams/Molar mass R = 8.31 L·kPa/ mol·K = 0.0821 L·atm/mol·K = 62.4 L·Torr/mol·K You must have a common set of units in the problem

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**Edward Jewitt Wheeler, Isaac Kaufman
Funk, William Seaver Woods, Arthur
Stimson Draper, Wilfred John Funk**

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web what type of cells are haploid 13 what is the haploid chromosome number in humans 14 how many autosomes are present in each human gamete how many sex chromosomes 15 complete the following table to summarize the differences between mitosis and meiosis use figure 6 2 to help you