

Chemistry 104 Final Exam With Completely Detailed Solutions 2024

How many moles of hydroxide ions are there in 25 mL of 0.010M solution of Ba(OH)₂?

- a. 7.5×10^{-1} mol
- b. 5.0×10^{-1} mol
- c. 2.5×10^{-1} mol
- d. 5.0×10^{-4} mol
- e. 2.5×10^{-4} mol - correct answer.d. 5.0×10^{-4} mol

What is the molarity of a solution that contains 46 grams of Na₂SO₄ (Mw = 142 g/mol) dissolved in enough water to make 750 mL of solution?

- a. 2.3M
- b. 0.60M
- c. 0.52M
- d. 0.43M
- e. 0.19M - correct answer.d. 0.43M

Imagine that you go into the lab and measure 35 mL of a 0.60M stock solution. You pour the 35 mL of stock solution into a beaker and dilute the solution by adding another 35 mL of water. What is the molarity of the new solution?

- a. 1.2M
- b. 0.60M
- c. 0.30M
- d. 0.15M
- e. 0.040M - correct answer.c. 0.30M

What is the correct expression for the equilibrium constant, K, for the following reaction?



If for a certain gas phase reaction mixture $Q = 1$, what can you say about the reaction?

- a. The reaction will proceed towards the product to achieve equilibrium.
- b. The equilibrium position will be determined by the value of K.
- c. The reaction will proceed to the left to achieve equilibrium.
- d. The reaction is already at equilibrium.

Jeff State Chemistry 104 Lab Manual Answers

M Lipman



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