

Ques. ► 6. $A = (n-1)d^1ns^1$, $B = (n-1)d^1ns^2$, Here $n = 4$
 [D.B.-16]

- What is common-ion effect?
- In which orbital between 3d and 4p electron will enter first?
- Determine the four quantum number of the last electron of the element A.
- Which among the elements A and B will produce colored compound? Explain with reason.

Answer to the question no. 6

1. The process by which the dissociation rate of the weak electrolyte is reduced due to the presence of a strong electrolyte is known as similar-ion effect.

2. In which orbital the electron will enter first is determined by Aufbau's principle or $(n + l)$. According to this principle, the orbital of which $(n + l)$ value is low has low energy and electron will enter to it first. But if the $(n + l)$ value of two orbitals are same, then electron will enter the orbital whose n value is lower. For both 3d and 4p orbital the $(n + l)$ value is 5. But for 3d orbital the n value is 3 and for 4p orbital the n value is 4. So, electron will enter 3d orbital first.

3. The last electron of the given element A enters into $(n - 1)d$ orbital.

As, the value of n is 4.

So, the value of four quantum numbers of the last electron are :

- Principal quantum number = $n - 1 = 4 - 1 = 3$
- The value of subsidiary quantum number, l for d orbital is 2.
- Magnetic quantum number, $m = 2$.

Ques. ► 7



At 20°C, K_{sp} of MQ = 3.5×10^{-4}
 K_{sp} of PN = 2.5×10^{-2}

[D.B.-16]

- Write down the law of mass action.
- Which between H_2SO_3 and HNO_3 is more acidic and why?
- If the MN solution of the stem is saturated then calculate the solubility product of MN at 20°C.
- If the two solutions in the stem are mixed, will there be any ppt? Explain with reason.

Answer to the question no. 7

1. At a fixed temperature the rate of any reaction is proportional to the active mass (partial pressure in case of liquid and concentration in case of gas) of its reactants. This is the law of mass action.

2. Between H_2SO_3 and HNO_3 , HNO_3 is more acidic. Because we know that the acid whose oxidation number of the central atom is more is more acidic.

Here for H_2SO_3 , the oxidation number of the central atom S is

Let the oxidation number of S be x .

$$2(+2) + x + 3(-2) = 0$$

$$4 + x - 6 = 0$$

$$x - 2 = 0$$

$$x = +2$$

For HNO_3 , the oxidation number of the central atom N is

$$1 + x + 3(-2) = 0$$

$$1 + x - 6 = 0$$

$$x - 5 = 0$$

$$x = +5$$

Since the oxidation number of N in HNO_3 is greater than that of S in H_2SO_3 , HNO_3 is more acidic.

3. If the MN solution of the stem is saturated then calculate the solubility product of MN at 20°C.

Let the solubility of MN be s mol/L.

$$K_{sp} = [M^{n+}][N^{m-}]$$

$$3.5 \times 10^{-4} = s \times s$$

$$s^2 = 3.5 \times 10^{-4}$$

$$s = \sqrt{3.5 \times 10^{-4}}$$

$$s = 1.87 \times 10^{-2} \text{ mol/L}$$

4. If the two solutions in the stem are mixed, will there be any ppt? Explain with reason.

When the two solutions are mixed, the concentration of M^{n+} and N^{m-} ions will be

$$[M^{n+}] = \frac{0.30 \text{ M} \times V_1}{V_1 + V_2}$$

$$[N^{m-}] = \frac{0.01 \text{ M} \times V_2}{V_1 + V_2}$$

Let $V_1 = V_2 = V$.

$$[M^{n+}] = \frac{0.30 \text{ M} \times V}{2V} = 0.15 \text{ M}$$

$$[N^{m-}] = \frac{0.01 \text{ M} \times V}{2V} = 0.005 \text{ M}$$

The ion product is

$$Q = [M^{n+}][N^{m-}] = 0.15 \times 0.005 = 7.5 \times 10^{-4}$$

Since $Q < K_{sp}$, no precipitate will form.

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Raffaela Di Napoli



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