

SECTION - I

- Q. 1. Select and write the most appropriate answers from the given alternatives for each sub-question :** [7]
- (i) In Van Arkel method of refining metal, impure zirconium is converted to unstable volatile compound by heating it with –
 (a) oxygen (b) chlorine (c) bromine (d) iodine
- (ii) The temperature at which vapour pressure of a liquid becomes equal to the atmospheric pressure is _____.
 (a) melting point (b) boiling point (c) 273 K (d) 373 K
- (iii) Which of the following parameters are correct for triclinic lattice ?
 (a) $\alpha = \beta = \gamma = 90^\circ$ and $a = b = c$ (b) $\alpha \neq \beta \neq \gamma \neq 90^\circ$ and $a \neq b \neq c$
 (c) $\alpha = \gamma = 90^\circ$, $\beta \neq 90^\circ$ and $a \neq b \neq c$ (d) $\alpha \neq \beta \neq \gamma \neq 90^\circ$ and $a \neq b \neq c$
- (iv) Which mixture is used for respiration by deep sea divers ?
 (a) He + O₂ (b) Ne + O₂ (c) Ar + O₂ (d) Kr + O₂
- (v) For the reaction $2\text{N}_2\text{O}_{5(l)} \longrightarrow 4\text{NO}_{2(l)} + \text{O}_{2(l)}$ in liquid bromine, which of the following rate equation is 'incorrect'
 (a) $-\frac{1}{2} \frac{d[\text{N}_2\text{O}_5]}{dt}$ (b) $-\frac{1}{4} \frac{d[\text{NO}_2]}{dt}$ (c) $\frac{d[\text{O}_2]}{dt}$ (d) $\frac{1}{4} \frac{d[\text{NO}_2]}{dt}$
- (vi) For a certain reaction $\Delta H = -50 \text{ KJ}$ and $\Delta S = -80 \text{ JK}^{-1}$, at what temperature does the reaction turn from spontaneous to nonspontaneous.
 (a) 6.25 K (b) 62.5 K (c) 625 K (d) 6250 K
- (vii) What is the ratio of volumes of H₂ and O₂ liberated during electrolysis of acidified water ?
 (a) 1 : 2 (b) 2 : 1 (c) 1 : 8 (d) 8 : 1
- Q. 2. Answer any THREE of the following :** [9]
- (i) Calculate ΔH° for the following reactions –
 $2\text{H}_3\text{BO}_{3(aq)} \longrightarrow \text{B}_2\text{O}_{3(s)} + 3\text{H}_2\text{O}_{(l)}$
 Given that
 (a) $\text{H}_3\text{BO}_{3(aq)} \longrightarrow \text{HBO}_{2(aq)} + \text{H}_2\text{O}_{(l)}$, $\Delta H^\circ_1 = -0.02 \text{ KJ}$
 (b) $\text{H}_2\text{B}_4\text{O}_{7(s)} \longrightarrow 2\text{B}_2\text{O}_{3(s)} + \text{H}_2\text{O}_{(l)}$, $\Delta H^\circ_2 = 17.3 \text{ KJ}$
 (c) $\text{H}_2\text{B}_4\text{O}_{7(s)} + \text{H}_2\text{O}_{(l)} \longrightarrow 4\text{HBO}_{2(aq)}$, $\Delta H^\circ_3 = -11.58 \text{ KJ}$
- (ii) Calculate molarity and molality of 6.3% solution of nitric acid having density 1.04 gcm^{-3} . (H = 1, N = 14, O = 16)
- (iii) What is the action of chlorine (Cl₂) on the following :
 (a) Cold and dilute caustic soda.
 (b) Hot and concentrated caustic soda.
 (c) Potassium bromide solution.
- (iv) Calculate the number of atoms present in 2 gram of crystal which has face-centred cubic (FCC) crystal lattice having edge length of 100 pm and density 10 gcm^{-3} .

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