

# HSC Chemistry Notes:

## Production of Materials:

### 1. Fossil Fuel products

#### 1.0. Construct Word and Balanced Formulae Equations of Chemical Equations that are encountered.



#### 1.1 Identify the Industrial Source for Ethylene from the cracking of some of the fractions from the refining of petroleum.

Oil refineries are required to balance their petroleum outputs to the demands of the marketplace, i.e. more gasoline produced for cars rather than grease. This doesn't occur in the natural process of fractional distillation and hence catalytic cracking is applied, converting lower demanded fractions into gasoline and ethylene (ethene).

**Catalytic Cracking:** High boiling point (molecular weight) fractions from crude oil are broken into lower boiling point (molecular weight) substances to increase the output of high demand products.



**Zeolites**, which are crystalline aluminosilicates, are used as a surface catalyst for catalytic cracking. Its ordered structure and large surface area allows it to effectively begin the 'cracking' process in its inner surface to outer, in 500 degrees plus and above atmospheric pressure, creating ethene and propene as by-products.

**Steam Cracking or Thermal cracking** is a non-catalyst cracking process where alkanes are passed extremely hot metal tubes, decomposing them into ethene, propene or butane.

#### 1.2 Identify that ethylene because of the high reactivity of its double bond is readily transformed into many useful products.

Most simple alkane and ethene are very similar in characteristics, they are both gases at room temperature and low boiling points, held together only by dispersion forces. However due to alkene's double bond in contrast to alkane's single bond, alkene's become more reactive by **opening out the double bond to form two single bonds, known as addition reactions.**

Alkenes such as ethene react with chlorine, bromine and iodine, decolorizing them.

**Note: Alkyl Groups** are alkane groups with one hydrogen missing.



Ethene can readily transform into useful products such as ethanol with the reaction of water and phosphoric acid as a catalyst.

Ethylene can also readily react with oxygen via a silver catalyst forming ethylene oxide (sterilizer) which then may be converted into ethylene glycol used as polymers and automotive antifreeze.

Can also be polymerized to polyethylene (plastic), and acetic acid (food preservative).

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**Atar**



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## **Table of Contents Hsc Chemistry Notes**

1. Understanding the eBook Hsc Chemistry Notes
  - The Rise of Digital Reading Hsc Chemistry Notes
  - Advantages of eBooks Over Traditional Books
2. Identifying Hsc Chemistry Notes
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Hsc Chemistry Notes
  - User-Friendly Interface
4. Exploring eBook Recommendations from Hsc Chemistry Notes
  - Personalized Recommendations
  - Hsc Chemistry Notes User Reviews and Ratings
  - Hsc Chemistry Notes and Bestseller Lists
5. Accessing Hsc Chemistry Notes Free and Paid eBooks
  - Hsc Chemistry Notes Public Domain eBooks
  - Hsc Chemistry Notes eBook Subscription Services
  - Hsc Chemistry Notes Budget-Friendly Options
6. Navigating Hsc Chemistry Notes eBook Formats

- ePub, PDF, MOBI, and More
  - Hsc Chemistry Notes Compatibility with Devices
  - Hsc Chemistry Notes Enhanced eBook Features
7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Hsc Chemistry Notes
    - Highlighting and Note-Taking Hsc Chemistry Notes
    - Interactive Elements Hsc Chemistry Notes
  8. Staying Engaged with Hsc Chemistry Notes
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Hsc Chemistry Notes
  9. Balancing eBooks and Physical Books Hsc Chemistry Notes
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Hsc Chemistry Notes
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Hsc Chemistry Notes
    - Setting Reading Goals Hsc Chemistry Notes
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Hsc Chemistry Notes
    - Fact-Checking eBook Content of Hsc Chemistry Notes
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development
    - Exploring Educational eBooks
  14. Embracing eBook Trends
    - Integration of Multimedia Elements
    - Interactive and Gamified eBooks

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