

## Question 2

- Which of the following salts has the least solubility in water at 25°C?

F  $\text{CaCO}_3$

G KBr

H CaCl

J KI

Ion	General Solubility Rule
$\text{NO}_3^-$	All nitrates are soluble
$\text{C}_2\text{H}_3\text{O}_2^-$	All acetates are soluble ( $\text{AgC}_2\text{H}_3\text{O}_2$ only moderately)
$\text{Cl}^-$ , $\text{Br}^-$ , $\text{I}^-$	All chlorides, bromides and iodides are soluble except $\text{Ag}^+$ , $\text{Pb}^+$ and $\text{Hg}_2^{2+}$ . ( $\text{PbCl}_2$ is slightly soluble in cold water and moderately soluble in hot water.)
$\text{SO}_4^{2-}$	All sulfates are soluble except those of $\text{Ba}^{2+}$ , $\text{Pb}^{2+}$ , $\text{Ca}^{2+}$ and $\text{Sr}^{2+}$
$\text{CO}_3^{2-}$ and $\text{PO}_4^{3-}$	All carbonates and phosphates are insoluble except those of $\text{Na}^+$ , $\text{K}^+$ and $\text{NH}_4^+$ . (Many acid phosphates are soluble).
$\text{OH}^-$	All hydroxides are insoluble except those of $\text{Na}^+$ and $\text{K}^+$ . Hydroxides of $\text{Ba}^{2+}$ and $\text{Ca}^{2+}$ are slightly soluble.
$\text{S}^{2-}$	All sulfides are insoluble except those of $\text{Na}^+$ , $\text{K}^+$ , $\text{NH}_4^+$ and those of the alkaline earths: $\text{Mg}^{2+}$ , $\text{Ca}^{2+}$ , $\text{Sr}^{2+}$ and $\text{Ba}^{2+}$ . (Sulfides of $\text{Al}^{3+}$ and $\text{Cr}^{3+}$ hydrolyze and precipitate as the corresponding hydroxides.)
$\text{Na}^+$ , $\text{K}^+$ and $\text{NH}_4^+$	All salts of sodium ion, potassium ion and ammonium ion are soluble except several uncommon ones.

# Lab Solubility Datasheet Answer Key

**G Orfield**



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