

$$PV = nRT$$

0.0°C 1 atm  
22.414 L

### Gas Stoichiometry Practice Sheet

- 1) For the reaction  $2 \text{H}_{2(g)} + \text{O}_{2(g)} \rightarrow 2 \text{H}_2\text{O}_{(l)}$ , how many liters of water can be made from 5 L of oxygen gas and an excess of hydrogen?

$$\frac{5 \cancel{\text{L O}_2}}{22.414 \cancel{\text{L O}_2}} \left| \frac{1 \text{ mol O}_2}{1 \text{ mol O}_2} \right| \left| \frac{2 \text{ mol H}_2\text{O}}{1 \text{ mol O}_2} \right| \left| \frac{22.414 \text{ L H}_2\text{O}}{1 \text{ mol H}_2\text{O}} \right| = 10 \text{ L H}_2\text{O}$$

0.0°C  
1 atm

- 2) How many liters of water can be made from 55 grams of oxygen gas and an excess of hydrogen at STP?

$$\frac{55 \cancel{\text{g O}_2}}{32 \cancel{\text{g O}_2}} \left| \frac{1 \text{ mol O}_2}{1 \text{ mol O}_2} \right| \left| \frac{2 \text{ mol H}_2\text{O}}{1 \text{ mol O}_2} \right| \left| \frac{22.414 \text{ L H}_2\text{O}}{1 \text{ mol H}_2\text{O}} \right| = 77 \text{ L H}_2\text{O}$$

- 3) How many liters of water can be made from 55 grams of oxygen gas and an excess of hydrogen at a pressure of 12.4 atm and a temperature of 85°C?

$$\text{O}_2 + 2\text{H}_2 \rightarrow 2\text{H}_2\text{O}$$

$$\frac{55}{\text{O}_2} \left| \frac{1 \text{ mol O}_2}{32.00 \text{ g O}_2} \right| \left| \frac{2 \text{ mol H}_2\text{O}}{1 \text{ mol O}_2} \right| = 3.4375 \text{ mol H}_2\text{O}$$

$$(12.4 \text{ atm}) V = (3.4375 \text{ mol H}_2\text{O}) \left( \frac{0.08205 \text{ L atm}}{\text{mol K}} \right) (358 \text{ K})$$

$$V = 8.14 \text{ L H}_2\text{O}$$

- 4) How many liters of water can be made from 34 grams of oxygen gas and

$$\frac{34}{\text{O}_2} \left| \frac{1 \text{ mol O}_2}{32 \text{ g O}_2} \right| \left| \frac{2 \text{ mol H}_2\text{O}}{1 \text{ mol O}_2} \right| \left| \frac{22.4 \text{ L H}_2\text{O}}{1 \text{ mol H}_2\text{O}} \right| = 47.8 \text{ L H}_2\text{O}$$

# General Chemistry Stoichiometry Test Answers

**JR Anderson**



**General Chemistry Stoichiometry Test Answers:**

## **General Chemistry Stoichiometry Test Answers** Book Review: Unveiling the Power of Words

In a global driven by information and connectivity, the ability of words has become more evident than ever. They have the capability to inspire, provoke, and ignite change. Such may be the essence of the book **General Chemistry Stoichiometry Test Answers**, a literary masterpiece that delves deep to the significance of words and their effect on our lives. Published by a renowned author, this captivating work takes readers on a transformative journey, unraveling the secrets and potential behind every word. In this review, we will explore the book's key themes, examine its writing style, and analyze its overall impact on readers.

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