

## Chem 1201 Final Exam

$$N_A = 6.022 \times 10^{23} \text{ things/mole}$$

$$1.6606 \times 10^{-24} \text{ g} = 1 \text{ amu}$$

---

- 1 Calculate the volume in liters of 82.0 m<sup>3</sup> of natural gas. 1 cm<sup>3</sup> = 1 mL.
- (A) 8.2 L  
(B)  $8.2 \times 10^7$  L  
(C)  $8.2 \times 10^4$  L  
(D) 0.082 L  
(E)  $8.2 \times 10^{-8}$  L
- 2 The molecule BrF<sub>3</sub> has a  electron domain geometry and a  molecular geometry.
- (A) tetrahedral, tetrahedral  
(B) trigonal bipyramidal, seesaw  
(C) tetrahedral, trigonal pyramidal  
(D) trigonal bipyramidal, T-shaped  
(E) trigonal planar, trigonal planar
- 3 What volume (mL) of 0.135 M sodium hydroxide is required to neutralize 440.0 mL of 0.584 M hydrochloric acid?
- (A)  $5.25 \times 10^{-3}$  mL  
(B) 34.7 mL  
(C)  $1.9 \times 10^3$  mL  
(D) 1.9 mL  
(E)  $6.47 \times 10^3$  mL
- 4 Which of these statements best describes the Heisenberg Uncertainty principle?
- (A) For each electron in an atom the magnetic quantum number  $m_l$  is uncertain.  
(B) The probability of finding an electron in an atom is determined by  $\psi^2$   
(C) In the atomic world all measurable quantities are completely uncertain.

# General Chemistry 1 Final Exam

**Jiyuan Zhang**



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