

Name \_\_\_\_\_

Chemistry Period \_\_\_\_\_

Worksheet Specific Heat

Specific Heat of water =  $4.18 \text{ J/g } ^\circ\text{C}$

1. A sample of mercury is heated from  $25.5^\circ\text{C}$  to  $52.5^\circ\text{C}$ . In the process,  $3050\text{J}$  of heat are absorbed. What mass of mercury was in the sample? The specific heat of mercury is  $0.140 \text{ J/g}^\circ\text{C}$ . **ANS: 807 g**
2. A block of aluminum with a mass of  $140\text{g}$  is cooled from  $98.4^\circ\text{C}$  to  $62.2^\circ\text{C}$  with a release of  $1137\text{J}$  of heat. From these data, calculate the specific heat of aluminum. **ANS: 0.224**
3. A total of  $226 \text{ J}$  of heat are absorbed as  $58.3 \text{ g}$  of lead is heated from  $12.0^\circ\text{C}$  to  $42.0^\circ\text{C}$ . From this data, what is the specific heat of lead? **ANS: 0.129**
4. What is the change in heat energy when  $132\text{g}$  of copper at  $32.2^\circ\text{C}$  is raised to  $45.0^\circ\text{C}$ ? The specific heat of copper is  $0.385 \text{ J/g } ^\circ\text{C}$ . **ANS: 650 J**
5. What is the change in heat energy when  $40.0\text{g}$  of water at  $25.0^\circ\text{C}$  is raised to  $60.0^\circ\text{C}$ ? **ANS: 5852 J**
6. What is the change in heat energy when  $75.0$  of water at  $9.8^\circ\text{C}$  is raised to  $22.4^\circ\text{C}$ ? **3950 J**
7. What is the change in heat energy when  $125\text{g}$  of water at  $10.0^\circ\text{C}$  is raised to  $90.0^\circ\text{C}$ ? **ANS: 41.8 kJ**
8. What is the change in heat energy when  $64.82 \text{ g}$  of aluminum metal at  $100.0^\circ\text{C}$  is cooled to  $82.0^\circ\text{C}$ ? The specific heat of aluminum is  $0.897 \text{ J/g } ^\circ\text{C}$ . **ANS: 1047**
9. What is the change in temperature if  $24\text{g}$  of water at  $50^\circ\text{C}$  absorbs  $985 \text{ J}$  of heat energy? **ANS: 9.82**
10. What is the mass of a sample of iron if that sample lost  $2300\text{J}$  of heat energy when it cooled from  $80^\circ\text{C}$  to  $30^\circ\text{C}$ ? The specific heat of iron is  $0.449 \text{ J/g } ^\circ\text{C}$ . **ANS: 102 g**
11. What is the specific heat of metal if its mass is  $26.86\text{g}$  and it requires  $418.6 \text{ J}$  of heat energy to raise its temperature from  $27.4^\circ\text{C}$  to  $67.3^\circ\text{C}$ ? **ANS: 0.391**
12. How much heat energy is needed to raise the temperature of  $59.7\text{g}$  of cadmium from  $25^\circ\text{C}$  to  $100^\circ\text{C}$ ? The specific heat of cadmium is  $0.231 \text{ J/g } ^\circ\text{C}$ . **ANS: 1034 J**

# Math Skills Specific Heat Answers

**Cutnell**



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web incoming kindergarteners summer learning packet dear margate families welcome to margate elementary we are so excited that you will be joining us in kindergarten please utilize this activity packet to help prepare your child for the upcoming year summer is an important time for each of us

[free printable worksheet bundles for summer learning](#) - Dec 27 2021

web apr 29 2022 whether you re on the hunt for practice tools for your students to work on at home activities for your summer school program or you re already thinking ahead to next year we ve got you covered we ve assembled free printable worksheet bundles for kindergarten through high school students