

Specific Heat Worksheet

$C = q/m\Delta T$, where q = heat energy, m = mass, and T = temperature Remember, $\Delta T = (T_{\text{final}} - T_{\text{initial}})$. Show all work and proper units.

1. A 15.75-g piece of iron absorbs 1086.75 joules of heat energy, and its temperature changes from 25°C to 175°C. Calculate the specific heat capacity of iron.

$$C = 1086.75 / 15.75 \cdot 150^\circ\text{C}$$

$$C = 0.46 \text{ J/g}^\circ\text{C}$$

2. How many joules of heat are needed to raise the temperature of 10.0 g of aluminum from 22°C to 55°C, if the specific heat of aluminum is 0.90 J/g°C?

$$0.90 \text{ J/g}^\circ\text{C} = q / 10.33^\circ\text{C}$$

$$0.90 = q / 330 \quad 0.90 \cdot 330 = q$$

$$q = 297 \text{ J}$$

3. To what temperature will a 50.0 g piece of glass raise if it absorbs 5275 joules of heat and its specific heat capacity is 0.50 J/g°C? The initial temperature of the glass is 20.0°C.

$$0.50 = (5275) / (50) \cdot (x - 20)$$

$$5275 = (0.50) (50) (x - 20)$$

$$5275 / 25 = 211$$

$$231^\circ\text{C}$$

$$= 211$$

$$211 = x - 20$$

$$+20 \quad +20$$

4. Calculate the heat capacity of a piece of wood if 1500.0 g of the wood absorbs 6.75×10^4 joules of heat, and its temperature changes from 32°C to 57°C.

5. 100.0 mL of 4.0°C water is heated until its temperature is 37°C. If the specific heat of water is 4.18 J/g°C, calculate the amount of heat energy needed to cause this rise in temperature.

6. 25.0 g of mercury is heated from 25°C to 155°C, and absorbs 455 joules of heat in the process. Calculate the specific heat capacity of mercury.

7. What is the specific heat capacity of silver metal if 55.00 g of the metal absorbs 47.3 calories of heat and the temperature rises 15.0°C?

8. If a sample of chloroform is initially at 25°C, what is its final temperature if 150.0 g of chloroform absorbs 1.0 kilojoules of heat, and the specific heat of chloroform is 0.96 J/g°C?

$$0.96 = (1000) / (150) \cdot (x - 25)$$

$$1000 = (0.96)(150)(x - 25)$$

$$1000 = 144(x - 25)$$

$$6.944 = x - 25$$

$$\frac{1.0 \text{ kJ} | 1000 \text{ J}}{1 \text{ kJ}}$$

$$x = 31.94^\circ\text{C} = 1000 \text{ J}$$

Holt Science Spectrum Answers Specific Heat

Christian G. Meyer



Holt Science Spectrum Answers Specific Heat:

Holt Science Spectrum Answers Specific Heat Book Review: Unveiling the Magic of Language

In an electronic era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its capability to stir emotions, provoke thought, and instigate transformation is truly remarkable. This extraordinary book, aptly titled "**Holt Science Spectrum Answers Specific Heat**," written by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound impact on our existence. Throughout this critique, we shall delve to the book is central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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