



6. Which of the images shown above depicts a phase that has definite volume but not definite shape?

- A) The one in the middle
 B) The one on the left
 C) The one in the middle and the one on the right
 D) The one on the right

7. Assuming all other parameters remain constant, what happens to the pressure of a gas when the volume of the gas is increased?

- A) increases
 B) decreases
 C) fluctuates sinusoidally
 D) remains the same

8. You collect 550 mL of argon gas at 23.0°C. What volume will the gas occupy at 46.0 °C if the pressure remains constant?

- A) 270 mL
 B) 532 mL
 C) 969 mL
 D) 1104 mL

9. What happens to the pressure of a gas inside a container if the temperature of the gas decreases?

- A) The pressure increases.
 B) The pressure does not change.
 C) The pressure decreases.
 D) The pressure cannot be predicted.

10. A sealed metal can full of air is heated from 20 °C to 120 °C. The can was initially at 1.0 atmosphere of pressure. The final pressure will be

- A) 6.0 atm
 B) 0.007 atm
 C) 0.72 atm
 D) 1.3 atm

11. The combined gas law relates which of the following?

- A) pressure and volume only
 B) temperature and pressure only
 C) volume and temperature only
 D) temperature, pressure, and volume

12. If I have 25 mL of a gas at a pressure of 2.1 atm and a temperature of 300 K, what will

the pressure be if I increase the temperature to 400 K and compress the gas to a volume

of 10 mL?

- A) 1.4 atm
 B) 8.6 atm
 C) 0.008 atm
 D) 7 atm
 E) none of these

13. A rigid container contains 2.50 mol O_2 , 0.90 mol N_2 , and 1.00 mol CO_2 . The total pressure of the system is 200 kPa. The partial pressure exerted by the N_2 in the mixture is

- A) 25.0 kPa
 B) 50.0 kPa
 C) 100 kPa
 D) 125 kPa
 E) 150 kPa

Gas Laws Test Answers

Hugh Stott Taylor



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