

Complete the following calculation by list the given information, rewriting the formula to solve for the unknown, and plugging in the information (number and units), and writing the answer significantly.

1. A sample of 0.500 moles of gas is placed in a container of volume of 2.50 L. What is the pressure of the gas in torr if the gas is at 25 °C? (Ideal Gas Law)

$$P = \frac{nRT}{V} = \frac{(0.500 \text{ moles}) (0.0821 \text{ L} \cdot \text{atm}) (298 \text{ K})}{(\text{mole} \cdot \text{K}) (2.50 \text{ L})} = P = 4.89 \text{ atm} = 3720 \text{ torr}$$

2. If 2.00 mol of gas occupies 4.50 L at STP. How much of the same gas will occupy 3.00 L at STP? (Avogadro's Law)

$$n_2 = \frac{V_2 n_1}{V_1} = \frac{(3.00 \text{ L}) (2.00 \text{ moles})}{(4.50 \text{ L})} = 1.33 \text{ moles}$$

3. Determine the partial pressures of each of the gases in the following mixture: 17.04 g NH<sub>3</sub>, 40.36 g Ne and 19.00 g F<sub>2</sub>. The gases are at 1.50 atm of pressure. (Dalton's Law of Partial Pressure & Mole Ratios)

$$\frac{17.04 \text{ g NH}_3}{17.04 \text{ g NH}_3} \left| \frac{1 \text{ mole NH}_3}{17.04 \text{ g NH}_3} \right| \left| \frac{1.000 \text{ mole NH}_3}{3.500 \text{ moles}} \right| \quad \frac{1.000 \text{ mole NH}_3}{3.500 \text{ moles}} = 0.2857 \times 1.50 \text{ atm} = 0.429 \text{ atm NH}_3$$

$$\frac{40.36 \text{ g Ne}}{20.18 \text{ g Ne}} \left| \frac{1 \text{ mole Ne}}{20.18 \text{ g Ne}} \right| \left| \frac{2.000 \text{ moles Ne}}{3.500 \text{ moles}} \right| \quad \frac{2.000 \text{ moles Ne}}{3.500 \text{ moles}} = 0.5714 \times 1.50 \text{ atm} = 0.857 \text{ atm Ne}$$

$$\frac{19.00 \text{ g F}_2}{38.00 \text{ g F}_2} \left| \frac{1 \text{ mole F}_2}{38.00 \text{ g F}_2} \right| \left| \frac{0.5000 \text{ moles F}_2}{3.500 \text{ moles}} \right| \quad \frac{0.5000 \text{ moles F}_2}{3.500 \text{ moles}} = 0.1429 \times 1.50 \text{ atm} = 0.214 \text{ atm F}_2$$

$$1.000 \text{ moles} + 2.000 \text{ moles} + 0.5000 \text{ moles} = 3.500 \text{ moles}$$

4. If at 1.00 atm of pressure water boils at 100. °C, at what temperature would water boil if the pressure is 600. torr? (This shows why food doesn't cook well at higher elevations) (Gay-Lussac's Law)

$$T_2 = \frac{T_1 P_2}{P_1} = \frac{(373 \text{ K}) (0.789 \text{ atm})}{(1.00 \text{ atm})} = 294 \text{ K}$$

5. Calculate the volume of 40.6 g of F<sub>2</sub> at STP. (Ideal Gas Law)

$$\frac{40.6 \text{ g F}_2}{38.00 \text{ g F}_2} \left| \frac{1 \text{ mole F}_2}{38.00 \text{ g F}_2} \right| \left| \frac{22.4 \text{ L F}_2}{1 \text{ mole F}_2} \right| \left| \frac{23.9 \text{ L F}_2}{22.4 \text{ L F}_2} \right|$$

or

$$\frac{40.6 \text{ g F}_2}{38.00 \text{ g F}_2} \left| \frac{1 \text{ mole F}_2}{38.00 \text{ g F}_2} \right| \left| \frac{1.07 \text{ mole F}_2}{1.07 \text{ mole F}_2} \right|$$

$$V = \frac{nRT}{P} = \frac{(1.07 \text{ moles}) (0.0821 \text{ L} \cdot \text{atm}) (273 \text{ K})}{(\text{mole} \cdot \text{K}) (1.00 \text{ atm})} = 24.0 \text{ L F}_2$$

6. If the winter pressure that most car tire tires should be at to wear evenly is 32.0 psi, what is the pressure in atm? (Pressure Conversion)

$$\frac{32.0 \text{ psi}}{14.7 \text{ psi}} \left| \frac{1.00 \text{ atm}}{14.7 \text{ psi}} \right| \left| \frac{2.18 \text{ atm}}{2.18 \text{ atm}} \right|$$

# Gas Laws Lab Answers

**Spencer L. Seager, Michael R.  
Slabaugh**



## Gas Laws Lab Answers:

## Reviewing **Gas Laws Lab Answers**: Unlocking the Spellbinding Force of Linguistics

In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is truly astonishing. Within the pages of "**Gas Laws Lab Answers**," an enthralling opus penned by a very acclaimed wordsmith, readers embark on an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve to the book is central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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