

Answers

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Chapter 1 The atomic nature of matter

1.1 Nanomaterials and nanoparticles

WE 1.1.1 2.43×10^{12} nm

- 8.35×10^7 nm
- a 1.35×10^7 nm b 4.2×10^6 nm
- 2500 times larger

Shape	Surface area	Volume	Surface area/Volume
Cube 2 cm \times 2 cm \times 2 cm	24 cm ²	8 cm ³	3
Sphere of radius 1.35 cm	23.9 cm ²	11 cm ³	2.17
Tube (cylinder) Radius of 1 cm Height of 2 cm	19 cm ²	6.3 cm ³	3

- a Sphere
c Sphere
d A large SA : V ratio is required for rapid cooling. A small SA : V ratio is required to minimise loss of heat when cold.
- There are two reasons why their size makes nanoparticles useful for transporting medicines. First, their small size makes it easy for the nanoparticles to get into the body by inhalation, injection or absorption through the skin. Second, their small size gives nanoparticles a large surface area compared to their volume. This means that a large number of molecules can be adsorbed onto the surface of a relatively small volume of nanoparticles.

1.2 The atomic world

- Noble gases
- a Element b Element c Compound
d Compound e Element f Compound
g Element h Compound
- An element is made up of just one type of atom and these atoms are identical. A compound is a molecule that contains at least two different types of atoms in a definite proportion.
- a Iron b Potassium c Tungsten d Lead

1.3 Inside atoms

- 10 000–100 000 larger
- Protons and neutrons found in the nucleus.
- The electrostatic attraction of the protons; the negative electrons are attracted to the positive protons and pulled towards them.

1.4 Classifying atoms

WE 1.4.1 92 protons, $238 - 92 = 143$ neutrons, 92 electrons

- Mass number
- $Z = 15$. Protons = 15, electrons = 15. Neutrons = $31 - 15 = 16$
- Nitrogen
- Isotopes of the same element have the same atomic number and therefore the same number of protons and electrons. The number of neutrons is different between isotopes of the same element; therefore they have different mass numbers.
- 2 neutrons.

1.5 Electronic structure of atoms

- Each line in an emission spectrum corresponds to a specific amount of energy. This energy is emitted when electrons from higher-energy electron shells transition to a lower-energy shell. Different lines indicate that there are differences in energy between shells. This is evidence that electrons are found in shells with discrete energy levels.
- Electrons revolve around the nucleus in fixed, circular orbits. Electrons' orbits correspond to specific energy levels in the atom. Electrons can only occupy fixed energy levels and cannot exist between two energy levels. Orbits of larger radii correspond to energy levels of higher energy. Light or electromagnetic radiation.

1.6 Electronic configuration and the shell model

WE 1.6.1 2, 8, 18, 6

WE 1.6.2 6 electrons

- a 2, 8 b 2, 8, 2 c 2, 8, 2 d 2, 8, 18, 7 e 2, 8, 2, 2
- a 2, 2 b 2, 8, 6 c 2, 8, 8 d 2, 8, 2
- a Helium, He b Fluorine, F c Aluminium, Al
d Nitrogen, N e Chlorine, Cl
- The electronic configuration is 2, 8, 18, 7. Therefore the number of valence electrons is 7.

1.7 The Schrödinger model of the atom

WE 1.7.1 $1s^2 2s^2 2p^6 3s^2 3p^6 3d^4 4s^2$

Element (atomic number)	Electronic configuration using the shell model	Electronic configuration using the subshell model
Boron (5)	2, 3	$1s^2 2s^2 2p^1$
Lithium (3)	2, 1	$1s^2 2s^1$
Chlorine (17)	2, 8, 7	$1s^2 2s^2 2p^6 3s^2 3p^5$
Sodium (11)	2, 8, 1	$1s^2 2s^2 2p^6 3s^1$
Neon (10)	2, 8	$1s^2 2s^2 2p^6$
Potassium (19)	2, 8, 8, 1	$1s^2 2s^2 2p^6 3s^2 3p^6 4s^1$
Scandium (21)	2, 8, 9, 2	$1s^2 2s^2 2p^6 3s^2 3p^6 3d^1 4s^2$
Iron (26)	2, 8, 14, 2	$1s^2 2s^2 2p^6 3s^2 3p^6 3d^6 4s^2$
Bromine (35)	2, 8, 18, 7	$1s^2 2s^2 2p^6 3s^2 3p^6 3d^{10} 4s^2 4p^5$

- The subshell model is a refinement of the shell model. The shell model proposed that all electrons in the one shell were of equal energy. Evidence from emission spectra indicated that there were different electronic energy levels (called subshells) within a shell.

Chapter 1 review

- a 5.0×10^7 nm b 1.2×10^7 nm c 2.0×10^{10} nm
- Zinc oxide nanoparticles are colourless.
- Atoms are hard, indivisible structures.
- Monoatomic = helium, neon, krypton. Molecules = sulfur, oxygen, nitrogen. Large network = copper, gold, diamond, tin.
- The nucleus.
- The protons and neutrons form the nucleus. The electrons are grouped in shells and occupy the space around the nucleus.
- The mass of a proton is approximately equal to the mass of a neutron and is about 1840 times the mass of an electron. The proton and electron have equal but opposite charges and the neutron has no charge.
- a Atomic number is 24; mass number is 52
b 24 electrons, 24 protons, $52 - 24 = 28$ neutrons.
- No. Isotopes have the same number of protons (atomic number) but different numbers of neutrons (and therefore different mass numbers). These atoms have different atomic numbers and different mass numbers.

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