Unit 8 Review

1. The balanced equation $P_4(s) + 6H_2(g) \rightarrow 4PH_3(g)$ tells us that 4.0 mol H_2
   - A) reacts with 2.0 mol P_4
   - B) produces 8.0 mol PH_3
   - C) cannot react with phosphorus
   - D) produces 2.7 mol PH_3
   - E) reacts with 4.0 mol P_4

3. A 1.3-mol sample of $KClO_3$ was decomposed according to the equation
   $2KClO_3(s) \rightarrow 2KCl(s) + 3O_2(g)$
   How many moles of O_2 are formed assuming 100% yield?
   - A) 0.9 mol
   - B) 1.1 mol
   - C) 1.3 mol
   - D) 0.7 mol
   - E) 2.0 mol

5. The rusting of iron is represented by the equation $4Fe + 3O_2 \rightarrow 2Fe_2O_3$. If you have a 1.58-mol sample of iron, how many moles of Fe_2O_3 will there be after the iron has rusted completely?
   - A) 0.527 mol
   - B) 0.790 mol
   - C) 1.05 mol
   - D) 1.58 mol
   - E) 2.37 mol

7. Consider the following reaction, where X represents an unknown element.
   $6X(s) + 2B_2O_3(s) \rightarrow B_6X_2(s) + 3XO_2(g)$
   If 175 g of X reacts completely with diboron trioxide to produce 2.43 mol of B_6X_2, what is the identity of X?
   - A) Ge
   - B) Mg
   - C) Si
   - D) Ni
   - E) C

9. Refer to the following equation:
   $4NH_3(g) + 7O_2(g) \rightarrow 4NO_2(g) + 6H_2O(g)$
   How many molecules of NO_2 are produced when 3.02 mol of ammonia is completely reacted?
   - A) 12.08
   - B) 3.64×10^{24}
   - C) 1.82×10^{24}
   - D) 139
   - E) none of these

11. In the reaction $N_2(g) + 3H_2(g) \rightarrow 2NH_3(g)$, how many moles of ammonia would be produced from 1.09 mol of hydrogen and excess nitrogen?
   - A) 1.42 mol
   - B) 3.27 mol
   - C) 0.727 mol
   - D) 2.18 mol
   - E) 0.363 mol

13. Refer to the following **unbalanced** equation: $C_6H_{14} + O_2 \rightarrow CO_2 + H_2O$
   What mass of oxygen (O_2) is required to react completely with 13.1 g of C_6H_{14}?
   - A) 3.98×10^{4} g
   - B) 23.1 g
   - C) 4.86 g
   - D) 46.2 g
   - E) 0.152 g

15. How many molecules of carbon dioxide will be formed if 7.57 g of propane is burned in the following reaction?
   $C_3H_8(g) + 5O_2(g) \rightarrow 3CO_2(g) + 4H_2O(g)$
   - A) 1.37×10^{23} molecules
   - B) 1.03×10^{23} molecules
   - C) 4.14×10^{23} molecules
   - D) 3.10×10^{23} molecules
   - E) 5.17×10^{23} molecules

18. Calculate the mass of carbon dioxide produced from 25.7 g of octane, C_8H_{18}, in the following reaction.
   $2C_8H_{18}(g) + 25O_2(g) \rightarrow 16CO_2(g) + 18H_2O(g)$
   - A) 79.2 g CO_2
   - B) 158 g CO_2
   - C) 89.1 g CO_2
   - D) 1.238 g CO_2
   - E) 1.800 g CO_2

19. Calculate the molecules of oxygen required to react with 19.2 g of sulfur in the following reaction.
   $2S(s) + 3O_2(g) \rightarrow 2SO_3(g)$
   - A) 1.03×10^{24} molecules O_2
   - B) 2.28×10^{23} molecules O_2
   - C) 1.71×10^{23} molecules O_2
   - D) 3.42×10^{23} molecules O_2
   - E) 5.13×10^{23} molecules O_2

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1. Understanding the eBook Multiple Choice Stoichiometry Test
    - The Rise of Digital Reading Multiple Choice Stoichiometry Test
    - Advantages of eBooks Over Traditional Books
2. Identifying Multiple Choice Stoichiometry Test
    - Exploring Different Genres
    - Considering Fiction vs. Non-Fiction
    - Determining Your Reading Goals
3. Choosing the Right eBook Platform
    - Popular eBook Platforms
    - Features to Look for in an Multiple Choice Stoichiometry Test
    - User-Friendly Interface
4. Exploring eBook Recommendations from Multiple Choice Stoichiometry Test
    - Personalized Recommendations
    - Multiple Choice Stoichiometry Test User Reviews and Ratings
    - Multiple Choice Stoichiometry Test and Bestseller Lists
5. Accessing Multiple Choice Stoichiometry Test Free and Paid eBooks

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- Multiple Choice Stoichiometry Test eBook Subscription Services
- Multiple Choice Stoichiometry Test Budget-Friendly Options

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   - Interactive Elements Multiple Choice Stoichiometry Test
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   - Joining Online Reading Communities
   - Participating in Virtual Book Clubs
   - Following Authors and Publishers Multiple Choice Stoichiometry Test
9. Balancing eBooks and Physical Books Multiple Choice Stoichiometry Test
   - Benefits of a Digital Library
   - Creating a Diverse Reading Collection Multiple Choice Stoichiometry Test
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    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
11. Cultivating a Reading Routine Multiple Choice Stoichiometry Test
    - Setting Reading Goals Multiple Choice Stoichiometry Test
    - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Multiple Choice Stoichiometry Test
    - Fact-Checking eBook Content of Multiple Choice Stoichiometry Test
    - Distinguishing Credible Sources
13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends
    - Integration of Multimedia Elements
    - Interactive and Gamified eBooks

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