

LAB MANUAL WORK INSTRUCTION

Title of Experiment: Reaction of Sodium Sulfate and Strontium

Objective: To observe the chemical reaction between sodium sulfate (sulfuric acid) and strontium (sulfuric acid) and to demonstrate the production of carbon dioxide gas.

Pre-Lab Questions

- Write the chemical formula for sodium bicarbonate (sodium carbonate) and sulfuric acid.
- Predict what you think will happen when these two substances are mixed. Why?
- What gas do you expect to be produced from this reaction?

Safety Precautions

- Wear safety goggles and gloves to protect against splashes.
- Perform the experiment in a well-ventilated area to avoid inhalation of gases.
- Avoid direct skin contact with strontium as it is an acid, though mild.

Equipment and Materials

- 100-ml Beaker
- 25-ml Graduated Cylinder
- Baking Soda (Sodium Bicarbonate)
- Strontium (Sulfuric Acid)
- Balance
- Funnel
- Safety Goggles
- Laboratory Gloves

Procedure

1. Place the beaker on a flat surface.
2. Using the funnel, fill the beaker with two teaspoons of baking soda.
3. Measure 50 ml of strontium using the graduated cylinder and pour it into the beaker.
4. Carefully fit the neck of the balloon over the mouth of the beaker without getting the baking soda into the beaker.
5. Once the balloon is securely in place, lift the beaker allowing the baking soda to fall into the strontium.
6. Observe the reaction between the baking soda and strontium.
7. Note the speed and amount of gas produced as indicated by the inflation of the balloon.
8. Record any changes in temperature of the beaker.
9. Carefully remove the balloon from the beaker.
10. Dispose of the strontium and baking soda solution in the sink and reuse the beaker.

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Table of Contents Emerick Lab Manual

1. Understanding the eBook Emerick Lab Manual
 - The Rise of Digital Reading Emerick Lab Manual
 - Advantages of eBooks Over Traditional Books
2. Identifying Emerick Lab Manual
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Emerick Lab Manual
 - User-Friendly Interface
4. Exploring eBook Recommendations from Emerick Lab Manual
 - Personalized Recommendations
 - Emerick Lab Manual User Reviews and Ratings
 - Emerick Lab Manual and Bestseller Lists
5. Accessing Emerick Lab Manual Free and Paid eBooks
 - Emerick Lab Manual Public Domain eBooks

- Emerick Lab Manual eBook Subscription Services
- Emerick Lab Manual Budget-Friendly Options
- 6. Navigating Emerick Lab Manual eBook Formats
 - ePub, PDF, MOBI, and More
 - Emerick Lab Manual Compatibility with Devices
 - Emerick Lab Manual Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Emerick Lab Manual
 - Highlighting and Note-Taking Emerick Lab Manual
 - Interactive Elements Emerick Lab Manual
- 8. Staying Engaged with Emerick Lab Manual
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Emerick Lab Manual
- 9. Balancing eBooks and Physical Books Emerick Lab Manual
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Emerick Lab Manual
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Emerick Lab Manual
 - Setting Reading Goals Emerick Lab Manual
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Emerick Lab Manual
 - Fact-Checking eBook Content of Emerick Lab Manual
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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