

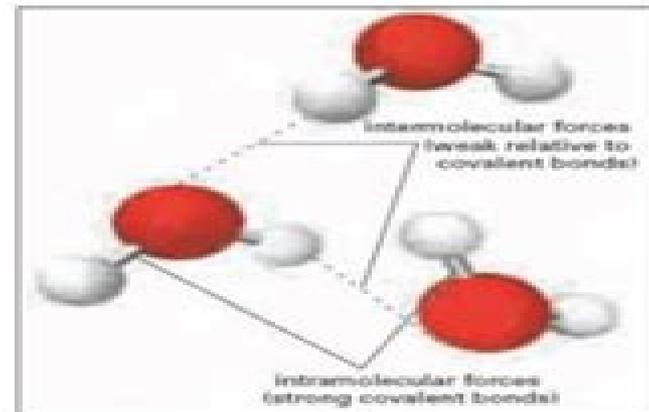
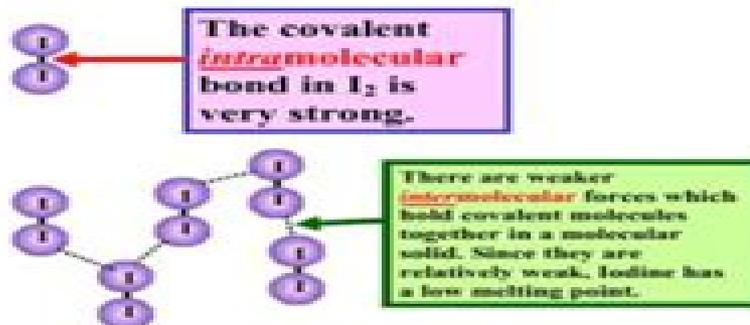
# Intermolecular Forces

## Intermolecular Forces:

- Forces between molecules
- Responsible for the state of matter: solid, liquid, or gas
- affect the melting and boiling points of compounds as well as the solubility of one substance in another
- weaker than covalent bonds within molecules, since molecular compounds melt easily (melting doesn't break the bond between atoms, but between the molecules themselves!)
  - o example: melting ice to water to steam will break apart the intermolecular forces and change state, but not decompose into  $H_2$  and  $O_2$

## Intramolecular Force (Force-Force attraction):

- Forces within molecules
- the attractive force between the valence electrons of one atom and the nucleus of another atom would be considered intermolecular forces
- Covalent bonds are examples of intramolecular forces



**Figure 3.20** Strong intramolecular forces (covalent bonds) hold the atoms in molecules together. Relatively weak intermolecular forces act between molecules.

# Exploring Intermolecular Forces Lab Answers

**Liying Dong**



## **Exploring Intermolecular Forces Lab Answers:**

## **Exploring Intermolecular Forces Lab Answers** Book Review: Unveiling the Magic of Language

In a digital era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its capability to stir emotions, provoke thought, and instigate transformation is actually remarkable. This extraordinary book, aptly titled "**Exploring Intermolecular Forces Lab Answers**," published by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we shall delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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