

$$4. a. \frac{10g \text{ Al} \mid 1 \text{ mol Al} \mid 3 \text{ mol Cu} \mid 63.55g \text{ Cu}}{26.98g \text{ Al} \mid 2 \text{ mol Al} \mid 1 \text{ mol Cu}} = 35.33g \text{ Cu}$$

$$\frac{20g \text{ CuCl}_2 \mid 1 \text{ mol CuCl}_2 \mid 3 \text{ mol Cu} \mid 63.55g \text{ Cu}}{134.45g \text{ CuCl}_2 \mid 3 \text{ mol CuCl}_2 \mid 1 \text{ mol Cu}} = \boxed{9.45g \text{ Cu}}$$

$$b. \frac{30g \text{ Al} \mid 1 \text{ mol Al} \mid 2 \text{ mol AlCl}_3 \mid 133.33g \text{ AlCl}_3}{26.98g \text{ Al} \mid 2 \text{ mol Al} \mid 1 \text{ mol AlCl}_3} = 148.25g \text{ AlCl}_3$$

$$\frac{100g \text{ CuCl}_2 \mid 1 \text{ mol CuCl}_2 \mid 2 \text{ mol AlCl}_3 \mid 133.33g \text{ AlCl}_3}{134.45g \text{ CuCl}_2 \mid 3 \text{ mol CuCl}_2 \mid 1 \text{ mol AlCl}_3} = \boxed{66.11g \text{ AlCl}_3}$$

$$c. \frac{5g \text{ Al} \mid 1 \text{ mol Al} \mid 3 \text{ mol Cu} \mid 63.55g \text{ Cu}}{26.98g \text{ Al} \mid 2 \text{ mol Al} \mid 1 \text{ mol Cu}} = 17.67g \text{ Cu}$$

~~10g / 17.67g x 100 = 56.6%~~

$$10g / 17.67g \times 100 = \boxed{56.6\%}$$

$$d. \frac{10g \text{ CuCl}_2 \mid 1 \text{ mol CuCl}_2 \mid 2 \text{ mol AlCl}_3 \mid 133.33g \text{ AlCl}_3}{134.45g \text{ CuCl}_2 \mid 3 \text{ mol CuCl}_2 \mid 1 \text{ mol AlCl}_3} = 6.61g \text{ AlCl}_3$$

~~5g / 6.61g x 100 = 75.6%~~

$$5g / 6.61g \times 100 = \boxed{75.6\%}$$

- 5.
- TRUE
 - FALSE
 - $\text{AgNO}_3 \Rightarrow 169.91g/\text{mol}$
 - 6.02×10^{23} particles = 1 mole of particles

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Arshad Iqbal



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