

Gas Law Stoichiometry Worksheet

Directions: Use significant figures and units in the problems below.

$$PV = nRT$$

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ALL
WORK!!!

1. Given the following unbalanced chemical equation for the combination reaction of sodium metal and chlorine gas:



- a. What volume of chlorine gas, measured at STP, is necessary for the complete reaction of 4.81 g of sodium metal.

$$4.81 \text{ g Na} \left| \frac{1 \text{ mol}}{23.0 \text{ g Na}} \right| \left| \frac{1 \text{ mol Cl}_2}{2 \text{ mol Na}} \right| \left| \frac{22.4 \text{ L}}{1 \text{ mol}} \right| = \boxed{2.34 \text{ L Cl}_2}$$

$$\boxed{2.34 \text{ L Cl}_2}$$



- a. Balance the above reaction.
b. What volume of oxygen gas at 25°C and 1.04 atm is needed for the complete combustion of 5.53 g of propane?

$$n = x$$

$$V = O_2(x)$$

$$T_1 = 25 + 273 = 298 \text{ K}$$

$$P_1 = 1.05 \text{ atm}$$

$$V = \left(\frac{5.53 \text{ g C}_3\text{H}_8}{44 \text{ g}} \right) \left(\frac{1 \text{ mol C}_3\text{H}_8}{1 \text{ mol C}_3\text{H}_8} \right) \left(\frac{5 \text{ mol O}_2}{1 \text{ mol C}_3\text{H}_8} \right) = 0.628 \text{ moles}$$

$$V = 14.6 \text{ L}$$

$$\boxed{14.6 \text{ L}}$$

3. Potassium permanganate, KMnO_4 , is produced commercially by oxidizing aqueous potassium manganate, K_2MnO_4 .



- a. Balance the above reaction.
b. What volume of $\text{Cl}_2(g)$, measured at STP, is needed to produce 10.0 g of KMnO_4 ?

$$V = \frac{nRT}{P}$$

$$10.0 \text{ g KMnO}_4 \left| \frac{1 \text{ mole}}{158 \text{ g}} \right| \left| \frac{1 \text{ mole Cl}_2}{2 \text{ mol}} \right| \left| \frac{22.4 \text{ L}}{1 \text{ mole Cl}_2} \right| = \boxed{0.709 \text{ L Cl}_2}$$

4. If water is added to magnesium nitride, ammonia gas is produced when the mixture is heated.



- a. Balance the above reaction.
b. If 10.3 g of magnesium nitride is treated with 10.3 g of water, what volume of ammonia gas would be collected at 24°C and 752 mmHg?

$$10.3 \text{ g Mg}_3\text{N}_2 \left| \frac{1 \text{ mole}}{100.9 \text{ g}} \right| \left| \frac{2 \text{ NH}_3}{1 \text{ mole}} \right| = 0.204 \text{ mol}$$

Lim Reagent

$$10.3 \text{ g H}_2\text{O} \left| \frac{1 \text{ mole}}{18 \text{ g}} \right| \left| \frac{2 \text{ NH}_3}{3 \text{ mole}} \right| = 0.381 \text{ mol}$$

$$24 + 273 = 297 \text{ K}$$

$$752 / 760 = 0.989 \text{ atm}$$

$$V = \frac{nRT}{P} = \frac{(0.204 \text{ mol} + 0.082 \text{ L} \cdot \text{atm} / \text{mol} \cdot \text{K}) (297 \text{ K})}{0.989 \text{ atm}} = \boxed{5.03 \text{ L}}$$

$$\boxed{5.03 \text{ L}}$$

$$V = \frac{nRT}{P}$$

Gas Stoichiometry Answer Key

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