

- 7) If I have 21 moles of gas held at a pressure of 78 atm and a temperature of 1000 K, what is the volume of the gas?

$$PV = nRT$$

$$V = \frac{nRT}{P} = \frac{21 \text{ mol} \cdot 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{mol} \cdot \text{K}} \cdot 1000 \text{ K}}{78 \text{ atm}}$$

$$V = 21.97 \text{ L}$$

- 8) If I have 1.9 moles of gas held at a pressure of 3 atm and in a container with a volume of 50 liters, what is the temperature of the gas?

$$PV = nRT$$

$$T = \frac{PV}{nR} = \frac{3 \text{ atm} \cdot 50 \text{ L}}{1.9 \text{ mol} \cdot 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{mol} \cdot \text{K}}}$$

$$T = 1602.07 \text{ K}$$

- 9) If I have 2.4 moles of gas held at a temperature of 47 °C and in a container with a volume of 47 liters, what is the pressure of the gas?

$$PV = nRT$$

$$P = \frac{nRT}{V} = \frac{2.4 \text{ mol} \cdot 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{mol} \cdot \text{K}} \cdot 320 \text{ K}}{47 \text{ L}}$$

$$P = 1.62 \text{ atm} \quad P = 169.64 \text{ kPa}$$

- 10) If I have an unknown quantity of gas held at a temperature of 1100 K in a container with a volume of 27 liters and a pressure of 780 atm, how many moles of gas do I have?

$$PV = nRT$$

$$n = \frac{PV}{RT} = \frac{780 \text{ atm} \cdot 27 \text{ L}}{0.0821 \frac{\text{L} \cdot \text{atm}}{\text{mol} \cdot \text{K}} \cdot 1100 \text{ K}}$$

$$n = 0.0076009 \text{ mol}$$

- 11) If I have 8.275 moles of gas at a temperature of 77 K and a pressure of 1.75 atmospheres, what is the volume of the gas?

$$PV = nRT$$

$$V = \frac{nRT}{P} = \frac{8.275 \text{ mol} \cdot 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{mol} \cdot \text{K}} \cdot 77 \text{ K}}{1.75 \text{ atm}}$$

$$V = 0.976 \text{ L}$$

- 12) If I have 72 liters of gas held at a pressure of 3.4 atm and a temperature of 225 K, how many moles of gas do I have?

$$PV = nRT$$

$$n = \frac{PV}{RT} = \frac{3.4 \text{ atm} \cdot 72 \text{ L}}{0.0821 \frac{\text{L} \cdot \text{atm}}{\text{mol} \cdot \text{K}} \cdot 225 \text{ K}}$$

$$n = 0.0796 \text{ mol}$$

Gas Laws Section 2 Review Answers

M Carnoy



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Table of Contents Gas Laws Section 2 Review Answers

1. Understanding the eBook Gas Laws Section 2 Review Answers
 - The Rise of Digital Reading Gas Laws Section 2 Review Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Gas Laws Section 2 Review Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gas Laws Section 2 Review Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Gas Laws Section 2 Review Answers
 - Personalized Recommendations

- Gas Laws Section 2 Review Answers User Reviews and Ratings
- Gas Laws Section 2 Review Answers and Bestseller Lists
- 5. Accessing Gas Laws Section 2 Review Answers Free and Paid eBooks
 - Gas Laws Section 2 Review Answers Public Domain eBooks
 - Gas Laws Section 2 Review Answers eBook Subscription Services
 - Gas Laws Section 2 Review Answers Budget-Friendly Options
- 6. Navigating Gas Laws Section 2 Review Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Gas Laws Section 2 Review Answers Compatibility with Devices
 - Gas Laws Section 2 Review Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Gas Laws Section 2 Review Answers
 - Highlighting and Note-Taking Gas Laws Section 2 Review Answers
 - Interactive Elements Gas Laws Section 2 Review Answers
- 8. Staying Engaged with Gas Laws Section 2 Review Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gas Laws Section 2 Review Answers
- 9. Balancing eBooks and Physical Books Gas Laws Section 2 Review Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gas Laws Section 2 Review Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Gas Laws Section 2 Review Answers
 - Setting Reading Goals Gas Laws Section 2 Review Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Gas Laws Section 2 Review Answers
 - Fact-Checking eBook Content of Gas Laws Section 2 Review Answers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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