

Gas Laws Practice

Equations and Constants

1 atm = 760 mm Hg

1 atm = 14.70 psi

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$$

$$PV = nRT$$

R = 0.0821 L·atm/mol·K

Avogadro's #: 6.02×10^{23}

Describing Gases

1. Standard atmospheric pressure is 1 atmosphere. Express standard atmospheric pressure in mm Hg, kilopascals, and bars.

The Gas Laws

2. A gas occupies a volume of 224 cm³ and a pressure of 1.51 bar. If the gas expands to a volume of 578 cm³, what will the new pressure be? (Assume the temperature remains constant.)
3. A cylinder with a constant volume of 2.80 L has a pressure of 32.0 psi at a temperature of 25.0 °C. If the cylinder is warmed to a temperature of 75.0 °C, what will be the pressure inside the cylinder?
4. A gas occupies a volume of 600.0 mL at a temperature of 25.00 °C. At what temperature would the gas occupy only half this volume?
5. In air-conditioning systems, compressed gases are allowed to expand, and this expansion results in cooling. A gas with a volume of 12.0 mL at a pressure of 8.0 bar at a temperature of 45 °C is allowed to expand to a volume of 40.0 mL at a pressure of 2.0 bar. What is the temperature of the gas after it expands?
6. What is the volume of 4.52 moles of gas, calculated at standard temperature and pressure?
7. What is the pressure of 12.5 moles of gas at a temperature of 360.0 K and a volume of 5.02 liters?
8. A helium balloon has a volume of 3.4 liters and a pressure of 1.05 atmospheres at 25 °C. How many moles of helium gas are in the balloon? How many grams of gas?

Diffusion and Effusion

9. What is the difference between diffusion and effusion?

Gas Laws Extra Practice

Albert Gibson, Robert McLean



Gas Laws Extra Practice:

Addison-Wesley Chemistry ,1990 *The Latest and Best of TESS* ,1991 **Basic Concepts of Chemistry** Leo J. Malone,Theodore Dolter,2008-12-03 Engineers who need to have a better understanding of chemistry will benefit from this accessible book It places a stronger emphasis on outcomes assessment which is the driving force for many of the new features Each section focuses on the development and assessment of one or two specific objectives Within each section a specific objective is included an anticipatory set to orient the reader content discussion from established authors and guided practice problems for relevant objectives These features are followed by a set of independent practice problems The expanded Making it Real feature showcases topics of current interest relating to the subject at hand such as chemical forensics and more medical related topics Numerous worked examples in the text now include Analysis and Synthesis sections which allow engineers to explore concepts in greater depth and discuss outside relevance CliffsStudySolver: Chemistry Charles Henrickson,2007-05-03 The CliffsStudySolver workbooks combine 20 percent review material with 80 percent practice problems and the answers to help make your lessons stick CliffsStudySolver Chemistry is for students who want to reinforce their knowledge with a learn by doing approach Inside you ll get the practice you need to learn Chemistry with problem solving tools such as Clear concise reviews of every topic Practice problems in every chapter with explanations and solutions A diagnostic pretest to assess your current skills A full length exam that adapts to your skill level A glossary examples of calculations and equations and situational tasks can help you practice and understand chemistry This workbook also covers measurement chemical reactions and equations and matter elements compounds and mixtures Explore other aspects of the language including Formulas and ionic compounds Gases and the gas laws Atoms The mole elements and compounds Solutions and solution concentrations Chemical bonding Acids bases and buffers Practice makes perfect and whether you re taking lessons or teaching yourself CliffsStudySolver guides can help you make the grade

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