

Mole Lab

Introduction to The Mole Concept



Introduction

Although technically not a laboratory experiment, this activity certainly helps to drive home the main idea behind the mole concept—that chemists can count out infinitesimally small particles by weighing.

Concepts

- Avogadro's number
- Chemical formulas
- Molar mass or molecular weight

Materials (for each lab station, two or three lab pairs can work at each station)

- Aluminum foil
- Ammonium sulfate, $(\text{NH}_4)_2\text{SO}_4$, 100 g
- Sodium chloride, NaCl , 500 g
- Water, distilled, 300 mL
- Balance, electronic, 0.01-g precision (if only triple beam balances are available, it is best if there is one for each lab pair)
- Beakers, 600-mL, or cut off 2-L soda bottles, 3 labeled
- Cups, blue, clear, and red plastic, 2–3 of each
- Pipets, Beral-type, 2–3
- Spoons, plastic, 4–6

Materials (for the instructor's station, preferably near the center of the room)

- Copper sheet, 12 pieces*
 - Balance, electronic, 0.01-g precision
 - Cups, blue, clear, and red plastic, 1 each
 - Scoring Standard Table
- *See the Preparation section below for size of copper pieces.*

Safety Precautions

Ammonium sulfate is slightly toxic by ingestion. Wear chemical splash goggles, chemical-resistant gloves, and a chemical-resistant apron. Wash hands thoroughly with soap and water before leaving the laboratory. Follow all laboratory safety guidelines. Please review current Material Safety Data Sheets for additional safety, handling, and disposal information.

Preparation

1. Label three 600-mL beakers “NaCl,” “water,” and “ammonium sulfate” and add to each the corresponding substance. *Note:* Do not write the chemical formula for ammonium sulfate on the beaker; students must determine this.
2. Place 2–3 spoons in each beaker containing a solid substance.
3. Place 2–3 pipets in the beaker of water.
4. Cut twelve pieces of copper sheet according to Table 1 and label with the corresponding numbers 1–12. *Note:* The copper pieces can have any weights you want them to. Using the weights given in the table allows you to use the Scoring Standard table as is. If you use other weights, just modify column C in the table accordingly, as shown in Figure 1 on page 2.

| Cu Piece | Mass (g) |
|----------|----------|
| #1 | 4.22 |
| #2 | 4.78 |
| #3 | 4.98 |
| #4 | 5.13 |
| #5 | 5.44 |
| #6 | 5.97 |
| #7 | 6.09 |
| #8 | 6.55 |
| #9 | 6.79 |
| #10 | 7.13 |
| #11 | 7.65 |
| #12 | 7.98 |

Mole Lab Student Guide

Valentine Korah



Mole Lab Student Guide:

The Organic Chem Lab Survival Manual James W. Zubrick, 2016-01-19 Written for the laboratory that accompanies the sophomore junior level courses in Organic Chemistry Zubrick provides students with a valuable guide to the basic techniques of the Organic Chemistry lab The book will help students understand and practice good lab safety It will also help students become familiar with basic instrumentation techniques and apparatus and help them master the latest techniques such as interpretation of infrared spectroscopy The guide is mostly macroscale in its orientation *Study Guide for Chemical Principles* Thomas Elliott Taylor, 1979

Student Study Guide and Solutions Manual to accompany General Organic and Biological Chemistry, 1e Kenneth W. Raymond, 2005-10-07 Finally readers have a shorter less intimidating introduction to general organic and biological chemistry Not only is Raymond's text concise it also takes an integrated approach to presenting important topics in a way that makes the material easier to understand In this approach similarities can be exploited and concepts reinforced The result is that readers see the strong connections that exist between these three branches of chemistry *General, Organic, and Biological Chemistry Study Guide and Selected Solutions* Karen C. Timberlake, 2001-11

Keyed to the learning goals in the text this guide is designed to promote active learning through a variety of exercises with answers and mastery exams The guide also contains complete solutions to odd numbered problems

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