

Naming Ionic Compounds

What are the structural units that make up ionic compounds and how are they named?

Why?

When working in chemistry, it is often convenient to write a chemical in symbols. For example we might write down the substance table salt as NaCl. In talking about chemistry however, it is a bit tacky to say "n-ay see-ell" when we want to refer to a substance. Also, in formal writing we should use the name of the compound rather than its symbols. Therefore we need to learn how to say the proper names of ionic substances.

Model 1 – Ion Charges for Selected Elements

H ⁺													
Li ⁺	Be ²⁺									N ³⁻	O ²⁻	F ¹⁻	
Na ⁺	Mg ²⁺	Transition elements								P ³⁻	S ²⁻	Cl ¹⁻	
K ⁺	Ca ²⁺	Fe ²⁺	Ni ²⁺	Cu ⁺	Zn ²⁺								Br ¹⁻
		Fe ³⁺	Ni ³⁺	Cu ²⁺									
Rb ⁺	Sr ²⁺			Ag ¹⁺						Sn ²⁺			I ¹⁻
										Sn ⁴⁺			
	Ba ²⁺									Pb ²⁺			
					Hg ₂ ²⁺					Pb ⁴⁺			
					Hg ²⁺								
← Cations →										← Anions →			

1. Based on the information in Model 1:

a. Identify three elements that form only one cation.

Cu⁺ K⁺ Na⁺

b. Identify three elements that form only one anion.

+ charged elements

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Lauren Gardner



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