

Gas Laws Worksheet - CP

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Use the Combined Gas Law (or its component formulas) to answer the following questions:

21) For questions (a → f) circle the answer that describes what would happen in each scenario:

- | | |
|--|---------------------------------|
| a) For Boyle's law if pressure increases, what happens to volume? | (Increase / Decrease / Neither) |
| b) For Boyle's Law, if volume decreases, what happens to pressure? | (Increase / Decrease / Neither) |
| c) For Charles's Law, if temperature increases, what happens to volume? | (Increase / Decrease / Neither) |
| d) For Charles's Law, if volume decreases, what happens to temperature? | (Increase / Decrease / Neither) |
| e) For Gay-Lussac's Law, if pressure increases, what happens to temperature? | (Increase / Decrease / Neither) |
| f) For Gay-Lussac's Law, if temperature decreases, what happens to pressure? | (Increase / Decrease / Neither) |

22) If 22.5 L of nitrogen at 748 mm Hg are compressed to 725 mm Hg at constant temperature. What is the new volume?

23) Calculate the decrease in temperature when 6.00 L at 20.0 °C is compressed to 4.00 L.

24) The gases in a hair spray can are at a temperature of 27 °C and a pressure of 3 atm . If the gases in the can reach a pressure of 9 atm, the can will explode. To what temperature must the gases be raised in order for the can to explode? Assume constant volume.

25) A gas balloon has a volume of 106.0 liters when the temperature is 45.0 °C and the pressure is 740.0 mm of mercury. What will its volume be at 20.0 °C and 780 .0 mm of mercury pressure?

26) What pressure is required to compress 196.0 liters of air at 1.00 atmosphere into a cylinder whose volume is 26.0 liters?

27) A gas occupies 900.0 mL at a temperature of 27.0 °C. What is the volume at 132.0 °C?

28) Maybeline Cousteau's backup oxygen tank reads 900 mmHg while on her boat, where the temperature is 27 °C. When she dives down to the bottom of an unexplored methane lake on a recently-discovered moon of Neptune, the temperature will drop down to -183 °C. What will the pressure in her backup tank be at that temperature?

29) If 10.0 liters of oxygen at STP are heated to 512 °C, what will be the new volume of gas if the pressure is also increased to 2.0 atm?

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RS Peters



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