

## Chapter 3 Mass Relationships in Chemical Reactions

Student: \_\_\_\_\_

- An atom of bromine has a mass about four times greater than that of an atom of neon. Which choice makes the correct comparison of the relative numbers of bromine and neon atoms in 1,000 g of each element?
  - The number of bromine and neon atoms is the same.
  - There are one thousand times as many bromine atoms as neon atoms.
  - There are one thousand times as many neon atoms as bromine atoms.
  - There are four times as many neon atoms as bromine atoms.
  - There are four times as many bromine atoms as neon atoms.
- An atom of bromine has a mass about four times greater than that of an atom of neon. How many grams of neon will contain the same number of atoms as 1,000 g of bromine?
  - 4 g Ne
  - 250 g Ne
  - 400 g Ne
  - 1,000 g Ne
  - 4,000 g Ne
- What is the average mass, in grams, of one atom of iron?
  - $6.02 \times 10^{-23}$  g
  - $1.66 \times 10^{-24}$  g
  - $9.28 \times 10^{-23}$  g
  - 55.85 g
  - $55.85 \times 10^{-23}$  g
- What is the mass, in grams, of one arsenic atom?
  - $5.48 \times 10^{-23}$  g
  - 33.0 g
  - 74.9 g
  - $1.24 \times 10^{-22}$  g
  - $8.04 \times 10^{-21}$  g
- What is the mass, in grams, of one copper atom?
  - $1.055 \times 10^{-22}$  g
  - 63.55 g
  - 1 amu
  - $1.66 \times 10^{-24}$  g
  - $9.476 \times 10^{-23}$  g

# Mass Relationships In Chemical Reactions Practice Test

**Ian Pickup**



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