

## Stoichiometry Worksheet #2 (mole-mass, mass-mole problems)



- a. If 15.0g of  $N_2O_4$  was produced, how many moles of  $O_2$  were required?

$$\frac{15.0\text{g } N_2O_4}{92.0\text{g } N_2O_4} \left| \frac{1 \text{ mol } N_2O_4}{1 \text{ mol } N_2O_4} \right| \frac{2 \text{ mol } O_2}{1 \text{ mol } N_2O_4} = 0.326 \text{ mol } O_2$$

- b. If  $4.0 \times 10^{-3}$  moles of oxygen reacted, how many grams of  $N_2$  were needed?

$$\frac{4.0 \times 10^{-3} \text{ mol } O_2}{2 \text{ mol } O_2} \left| \frac{1 \text{ mol } N_2}{2 \text{ mol } O_2} \right| \frac{28.0\text{g } N_2}{1 \text{ mol } N_2} = 5.6 \times 10^{-2} \text{ g } N_2$$



- a. What is the mass of potassium nitrate that is produced when 2.04 moles of potassium phosphate react?

$$\frac{2.04 \text{ mol } K_3PO_4}{1 \text{ mol } K_3PO_4} \left| \frac{3 \text{ mol } KNO_3}{1 \text{ mol } K_3PO_4} \right| \frac{101.1 \text{ g } KNO_3}{1 \text{ mol } KNO_3} = 619\text{g } KNO_3$$

- b. If 5.80g of aluminum phosphate are formed, how many moles of aluminum nitrate reacted?

$$\frac{5.80\text{g } AlPO_4}{122.0\text{g } AlPO_4} \left| \frac{1 \text{ mol } AlPO_4}{122.0\text{g } AlPO_4} \right| \frac{1 \text{ mol } Al(NO_3)_3}{1 \text{ mol } AlPO_4} = 4.75 \times 10^{-2} \text{ mol } Al(NO_3)_3$$



- a. If you have 5.50 mol of  $CaC_2$ , how much  $C_2H_2$  do you get?

$$\frac{5.50 \text{ mol } CaC_2}{1 \text{ mol } CaC_2} \left| \frac{1 \text{ mol } C_2H_2}{1 \text{ mol } CaC_2} \right| \frac{26.0\text{g } C_2H_2}{1 \text{ mol } C_2H_2} = 143\text{g } C_2H_2$$

- b. How many moles of water are needed when 65.0g of  $CaC_2$  have reacted?

$$\frac{65.0\text{g } CaC_2}{64.1\text{g } CaC_2} \left| \frac{1 \text{ mol } CaC_2}{64.1\text{g } CaC_2} \right| \frac{2 \text{ mol } H_2O}{1 \text{ mol } CaC_2} = 2.03 \text{ mol } H_2O$$

4. In photosynthesis, water reacts with carbon dioxide to give oxygen and glucose ( $C_6H_{12}O_6$ ). Write and balance the chemical equation. How many moles of  $CO_2$  are required to make 120.0g of glucose?



$$\frac{120.0\text{g } C_6H_{12}O_6}{180.0 \text{ g } C_6H_{12}O_6} \left| \frac{1 \text{ mol } C_6H_{12}O_6}{180.0 \text{ g } C_6H_{12}O_6} \right| \frac{6 \text{ mol } CO_2}{1 \text{ mol } C_6H_{12}O_6} = 4.000 \text{ mol } CO_2$$

# Mass Mass Stoichiometry Packet Answers

**Anthony S. Fauci**



**Mass Mass Stoichiometry Packet Answers:**

## Reviewing **Mass Mass Stoichiometry Packet Answers**: Unlocking the Spellbinding Force of Linguistics

In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is truly astonishing. Within the pages of "**Mass Mass Stoichiometry Packet Answers**," an enthralling opus penned by a very acclaimed wordsmith, readers attempt an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve to the book is central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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