

- Structure, Hybridization, Resonance, Aromaticity
 - # of bonds- 1= sp (e.g. 3 bonds = sp^2)
 - Aromaticity= $4n+2$ electrons
 - Strained ring causes greater heat of combustion
- Acids and Bases
 - Lewis acid= accepts electron pair
 - Higher percentage of s character means more acidic hydrogen
- Stereoisomerism
 - Constitutional Isomers
 - Isomers that have different atom connectivity
 - Stereoisomers
 - Same connectivity but different arrangement in space
 - Enantiomers
 - Stereoisomers that are nonsuperimposable mirror reflections
 - Diastereomers
 - Stereoisomers that are not mirror reflections of each other
 - Stereocenter= sp^3 carbon bonded to four different substituents
 - Assigning R/S configuration
 - If lowest group is pointing away, determine normally
 - If lowest group is not pointing away, exchange with group pointing away and determine reverse

- Nucleophilic Substitutions and Eliminations

| ◦ Alkyl Halide | ◦ Strong Base | ◦ Weak Base | ◦ Poor Nucleophile |
|----------------|---------------|--------------------------------------|---------------------------------------|
| ◦ 1 | ◦ S_N2 | ◦ S_N2 | ◦ No rxn. |
| ◦ 2 | ◦ Mostly E2 | ◦ S_N2 | ◦ $S_N1/E1$ in polar, protic solvents |
| ◦ 3 | ◦ E2 | ◦ $S_N1/E1$ in polar, protic solvent | ◦ $S_N1/E1$ in polar, protic solvents |

- Relative strength of nucleophiles
 - Negatively charged nucleophiles react faster in substitution reactions than neutral nucleophiles
 - Nucleophilic strength is in the same order as base strength when the nucleophilic atoms are in the same period of the periodic table
 - Nucleophilic strength increases with increasing atom size when the nucleophilic atoms are in the same column of the periodic table
- Leaving group ability
 - Increases with decreasing base strength
- Base strength and structure

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Wolfgang Guggemos



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there has been a change of norm in the society of states in the 1990s the book defends the broad based constructivist claim
that state actions will be constrained if they cannot be legitimated and that new norms enable new practices but do not
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