

## Released AP Multiple Choice Questions – Acids &amp; Bases / Aqueous Equilibrium

D	33. The pH of 0.1-molar ammonia is approximately (A) 1 (B) 4 (C) 7 (D) 11 (E) 14
E	48. Which of the following ions is the strongest Lewis acid? (A) $\text{Na}^+$ (B) $\text{Cl}^-$ (C) $\text{CH}_3\text{COO}^-$ (D) $\text{Mg}^{2+}$ (E) $\text{Al}^{3+}$
C	49. Each of the following can act as both a Brønsted acid and a Brønsted base EXCEPT (A) $\text{HCO}_3^-$ (B) $\text{H}_2\text{PO}_4^-$ (C) $\text{NH}_4^+$ (D) $\text{H}_2\text{O}$ (E) $\text{HS}^-$
B	53. Which, if any, of the following species is in the greatest concentration in a 0.100-molar solution of $\text{H}_2\text{SO}_4$ in water? (A) $\text{H}_2\text{SO}_4$ molecules (B) $\text{H}_3\text{O}^+$ ions (C) $\text{HSO}_4^-$ ions (D) $\text{SO}_4^{2-}$ ions (E) All species are in equilibrium and therefore have the same concentrations.
E	71. Which of the following reactions does NOT proceed significantly to the right in aqueous solutions? (A) $\text{H}_3\text{O}^+ + \text{OH}^- \rightarrow 2 \text{H}_2\text{O}$ (B) $\text{HCN} + \text{OH}^- \rightarrow \text{H}_2\text{O} + \text{CN}^-$ (C) $\text{Cu}(\text{H}_2\text{O})_6^{2+} + 4 \text{NH}_3 \rightarrow \text{Cu}(\text{NH}_3)_4^{2+} + 4\text{H}_2\text{O}$ (D) $\text{H}_2\text{SO}_4 + \text{H}_2\text{O} \rightarrow \text{H}_3\text{O}^+ + \text{HSO}_4^-$ (E) $\text{H}_2\text{O} + \text{HSO}_4^- \rightarrow \text{H}_2\text{SO}_4 + \text{OH}^-$
E	75. If the acid dissociation constant, $K_a$ , for an acid HA is $8 \times 10^{-4}$ at $25^\circ\text{C}$ , what percent of the acid is dissociated in a 0.50-molar solution of HA at $25^\circ\text{C}$ ? (A) 0.08% (B) 0.2% (C) 1% (D) 2% (E) 4%
A	34. All of the following species can function as Brønsted-Lowry bases in solution EXCEPT (A) $\text{H}_2\text{O}$ (B) $\text{NH}_3$ (C) $\text{S}^{2-}$ (D) $\text{NH}_4^+$ (E) $\text{HCO}_3^-$

# General Equilibrium Ap Chemistry Multiple Choice

**AW Rasmussen**

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## **Table of Contents General Equilibrium Ap Chemistry Multiple Choice**

1. Understanding the eBook General Equilibrium Ap Chemistry Multiple Choice
  - The Rise of Digital Reading General Equilibrium Ap Chemistry Multiple Choice
  - Advantages of eBooks Over Traditional Books
2. Identifying General Equilibrium Ap Chemistry Multiple Choice
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an General Equilibrium Ap Chemistry Multiple Choice
  - User-Friendly Interface
4. Exploring eBook Recommendations from General Equilibrium Ap Chemistry Multiple Choice
  - Personalized Recommendations
  - General Equilibrium Ap Chemistry Multiple Choice User Reviews and Ratings
  - General Equilibrium Ap Chemistry Multiple Choice and Bestseller Lists
5. Accessing General Equilibrium Ap Chemistry Multiple Choice Free and Paid eBooks

- General Equilibrium Ap Chemistry Multiple Choice Public Domain eBooks
  - General Equilibrium Ap Chemistry Multiple Choice eBook Subscription Services
  - General Equilibrium Ap Chemistry Multiple Choice Budget-Friendly Options
6. Navigating General Equilibrium Ap Chemistry Multiple Choice eBook Formats
    - ePub, PDF, MOBI, and More
    - General Equilibrium Ap Chemistry Multiple Choice Compatibility with Devices
    - General Equilibrium Ap Chemistry Multiple Choice Enhanced eBook Features
  7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of General Equilibrium Ap Chemistry Multiple Choice
    - Highlighting and Note-Taking General Equilibrium Ap Chemistry Multiple Choice
    - Interactive Elements General Equilibrium Ap Chemistry Multiple Choice
  8. Staying Engaged with General Equilibrium Ap Chemistry Multiple Choice
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers General Equilibrium Ap Chemistry Multiple Choice
  9. Balancing eBooks and Physical Books General Equilibrium Ap Chemistry Multiple Choice
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection General Equilibrium Ap Chemistry Multiple Choice
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine General Equilibrium Ap Chemistry Multiple Choice
    - Setting Reading Goals General Equilibrium Ap Chemistry Multiple Choice
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of General Equilibrium Ap Chemistry Multiple Choice
    - Fact-Checking eBook Content of General Equilibrium Ap Chemistry Multiple Choice
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development

- Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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