

Standardized Test prep Answers

Chapter 1 Page 25

1. C                      4. B                      7. C  
 2. C                      5. C                      8. A  
 3. D                      6. C                      9. B

10. Answers may vary.

11. Observations C and G are chemical properties; the others are physical properties.

12. In a chemical change, one or more substances are converted into different substances. A physical change does not involve a change in the identity of the substance or substances present.

13. metals: shiny; good conductors of heat; good conductors of electricity; malleable or ductile; most are solids at room temperature  
 nonmetals: poor conductors of heat; poor conductors of electricity; many are gases at room temperature; those that are solids are brittle rather than malleable or ductile metalloids: properties intermediate between those of metals and nonmetals; less malleable than metals but not as brittle as solid nonmetals; most are semiconductors of electricity

Chapter 3 Page 93

1. C                      4. A                      7. B  
 2. D                      5. B                      8. D  
 3. C                      6. A                      9. C

10. Argon-40 has 22 neutrons ( $40 - 18 = 22$ ), and potassium-40 has 21 neutrons ( $40 - 19 = 21$ ).

11. 6.17 g

12. All cathode rays are the same, regardless of their source. Therefore, the particles responsible for the cathode rays must be present in all atoms. The particles are electrons.

13. When the average atomic mass is calculated, it is 10.811. Because the atomic mass is the same as the atomic mass of boron, mythium was not a new element.

Chapter 5 Page 1

1. C  
 2. A  
 3. A  
 4. A  
 5. C  
 6. A  
 7. D  
 8. B  
 9. C

10. S: 2251 kJ/mol; Cl: 2297 kJ/mol; Ar: 2666 kJ/mol; K: 3051 kJ/mol; Ca: 1345 kJ/mol; Sc: 1235 kJ/mol; Ti: 1310 kJ/mol. For the second ionization, the general trend is for increasing IE<sub>2</sub> across the period in Groups 2–18 with Group 1 having the highest IE<sub>2</sub>. IE<sub>2</sub> decreases going down a group.

11. Group 16 most commonly forms 2- ions, because these elements require only two more electrons to fill their shell (obtain a noble-gas configuration).

Chapter 2 Page 63

1. B                      4. B                      7. A  
 2. C                      5. A                      8. B  
 3. C                      6. B                      9. B

10. 2 cups; direct proportion

11. Statements A and B contain exact numbers.

12. The type of fertilizer is the variable being tested. Control factors are the types of radishes, the amount of water and the amount of sunshine. One control row should be planted under the same control factors but with no fertilizer. There are at least four things that could be used to determine the results: size, quantity, appearance, and taste. Analysis might include bar graphs of each of these measurements for each of the five fertilizer types and the no-fertilizer control row.

13. A unit must be defined in a way that does not depend on the circumstances of the measurement. Not every thumbnail is the same size.

Chap 4 Page 129

1. A                      4. D                      7. C  
 2. D                      5. B                      8. B  
 3. A                      6. B                      9. C

10. The color will be orange. Converting energy into frequency gives  $4.8 \times 10^{14}$ , which corresponds to the frequency of orange light.

11.  $\frac{1}{3\lambda}$      $\frac{1}{3\lambda}$      $\frac{1}{3\lambda}$      $\frac{1}{3\lambda}$

12. Electrons in atoms can occupy orbitals of only specific energies. When an atom is excited, the electron is no longer in the ground state. When the electron returns to a lower energy level, light is emitted. Because only specific energies are allowed, certain wavelengths of light are emitted, giving rise to the individual lines in the spectrum.

13. Photons of blue light are higher energy than photons of yellow light. Electrons can be emitted only when a photon of sufficient energy strikes the surface of the metal. Therefore, the energy of blue light is greater than the threshold energy, but the energy of yellow light is not.

Chapter 6 Page 215

1. A                      2. C                      7. A  
 3. B                      4. D                      8. B  
 5. A                      6. D                      9. C

10. Hybridization explains how the orbitals of an atom become rearranged when the atom forms covalent bonds.

11. Ionic crystals are brittle because shifting of the layers of ions results in large repulsive forces that cause the layers to part completely.

12. Naphthalene has the stronger intermolecular forces even though it is non-polar, because its boiling point is higher than that of acetic acid. Boiling point is directly correlated to strength of intermolecular forces; the stronger the intermolecular forces, the more energy needed to break all the intermolecular forces, and therefore the higher the boiling point. Naphthalene is so large that its dispersion forces are greater than the sum of the dispersion forces and hydrogen bonding in acetic acid.

13. There is a repulsive region at small bond distances, a minimum well at an intermediate bond distance, and an asymptotic region at large bond distances. The repulsive region is due to the repulsion between the two positive nuclei at distances where the nuclei are close to each other. The asymptotic region is where the atoms are far apart and therefore do not interact with each other. The potential energy decreases as the atoms come closer together and are attracted to each other. The minimum well is the region where the attractive forces are balanced by the repulsive forces. This distance is the equilibrium bond length.

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web jun 10 2022 in this video i will show you how to make your own ash glaze and it s actually a very special ash glaze because i will be using the left over ashes i get from pitfire i will show you

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web apr 7 2012 base ingredients 31 whiting 24 kentucky ball clay 22 5 silica 15 oak wood ash 5 custer feldspar 2 5 dolomite add the below colorants to base glaze 8 red iron oxide 3 8 cobalt carbonate 0 25 cobalt carbonate 1 iron chromate or 0 5 red iron oxide 0 5 chromium oxide 5 copper carbonate 3 tin oxide 1 6 copper

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web ash glazes ash glazes are among the oldest glazes in the ceramics business shigaraki and bizen ceramics from japan are famous all over the world due to the wood fired kiln ash was always available in a good cleanness and amounts ash glazes in general have a very simple composition

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web ash glaze traditionally important in east asia simply made from wood or plant ash which contains potash and lime feldspathic glazes of porcelain lead glazes plain or coloured are glossy and transparent after firing which need only about 800 c 1 470 f

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web fully updated and revised with new photographs and glaze recipes this is the third edition of this classic guide to ash glazes forever curious and eager to learn new things about ceramics phil rogers constantly tinkered with clay bodies glaze formulae and approaches to firing this volume is his seminal work on transforming ash into glaze an essential

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web ash glazes are ceramic glazes made from the ash of various kinds of wood or straw they have historically been important in east asia especially chinese pottery korean pottery and japanese pottery 1 many traditionalist east asian potteries still use ash glazing and it has seen a large revival in studio pottery in the west and east

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web oct 9 2023 a glaze is the glassy surface that one sees on most pottery whether commercial or that of an individual

potter artist it is one of the fundamental elements in ceramics and ash glazes in particular have long been favoured in asia  
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web bradley charles birkhimer my thesis is a documentation of my research with wood ash glaze specifics of the type of wood ash used how it was prepared where it was collected and glaze formulas are discussed i have included a brief history of ash glazes and the importance through their elemental structure

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web wadding a mixture of half epk kaolin and half alumina hydrate used to elevate ceramics off of shelves during atmospheric firings wood ash the mineral byproduct of burning wood which is primarily whiting silica and metal alkaline and alkaline earth oxides wood firing the process of firing ceramics solely with wood such that the ash

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web ash glazes have been used for over 2000 years potters who stoked their long burning kilns with wood noticed that after many firings the interior kiln bricks glistened the ash from the wood was coating the bricks adhering and melting as a thin layer of glaze

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web ash glazes examines the practicalities of collecting and testing wood ashes demonstrates the process of making them into glazes and offers a step by step guide to using them to decorate your pots this edition updated and revised by hajeong lee rogers is a celebration of pottery at its best

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round table the topic for this episode is ash glazes ash glazes

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web feb 12 2003 ash glazes phil rogers university of pennsylvania press feb 12 2003 art 176 pages ash glaze as the name suggests is a glaze derived from ashes