

AP[®] Chemistry GASES

The gaseous state of matter is the simplest and best-understood state of matter. You inhale approximately 8,500 L of air each day. This amounts to about 25 lbs of air. Breathing is a three-step process: inhaling, gas exchange with the circulatory system, and exhaling. Approximately 80% pressure differences created by your body allow you to breathe. Nearly the Earth's entire atmosphere is made up of only five gases: nitrogen, oxygen, water vapor, argon, and carbon dioxide.

PROPERTIES OF GASES

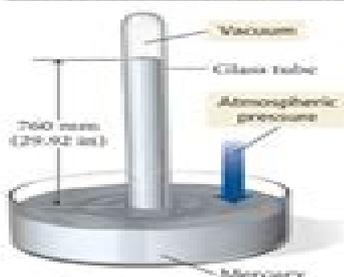
Only **FOUR** quantities are needed to define the state of a gas:

1. the quantity of the gas, n (in moles)
2. the temperature of the gas, T (in **KELVINS**)
3. the volume of the gas, V (in liters)
4. the pressure of the gas, P (in atmospheres)

A gas uniformly fills any container, is easily compressed & mixes completely with any other gas.



The Mercury Barometer



Pressure and Density



GAS PRESSURE is a measure of the force that a gas exerts on its container. Force is the physical quantity that interferes with inertia. Gravity is the force responsible for weight.

Force = mass \times acceleration; (Newton's 2nd Law)

The SI units follow:

$$N = kg \times m/s^2$$

Pressure-- Force/ unit area; N/m^2 ; which is the definition of 1.0 Pascal

Barometer--invented by Evangelista Torricelli in 1643; uses the height of a column of mercury to measure gas pressure (especially atmospheric P)

$$1 \text{ mm of Hg} = 1 \text{ torr}$$

$$1.00 \text{ atm} = 760.00 \text{ mm Hg} = 760.00 \text{ torr} = 29.92 \text{ in Hg} = 14.7 \text{ psi} = 101.325 \text{ kPa} \approx 10^5 \text{ Pa}$$

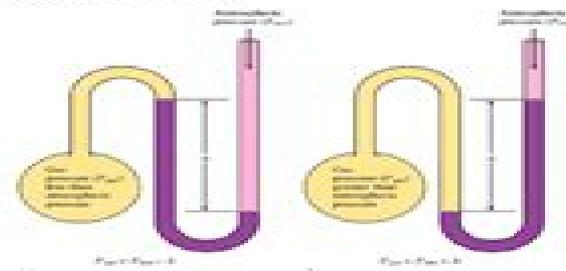
At sea level all of the above define **STANDARD PRESSURE**.

The SI unit of pressure is the Pascal (named after Blaise Pascal);

$$1 \text{ Pa} = 1 \text{ N} / \text{m}^2$$

The **manometer**—a device for measuring the pressure of a gas in a container. The pressure of the gas is given by h [the difference in mercury levels] in units of torr (equivalent to mm Hg).

- a) Gas pressure = atmospheric pressure - h
- b) Gas pressure = atmospheric pressure + h



Gases Packet In Chemistry

R Barnett



Gases Packet In Chemistry:

Decoding **Gases Packet In Chemistry**: Revealing the Captivating Potential of Verbal Expression

In a time characterized by interconnectedness and an insatiable thirst for knowledge, the captivating potential of verbal expression has emerged as a formidable force. Its ability to evoke sentiments, stimulate introspection, and incite profound transformations is genuinely awe-inspiring. Within the pages of "**Gases Packet In Chemistry**," a mesmerizing literary creation penned by way of a celebrated wordsmith, readers attempt an enlightening odyssey, unraveling the intricate significance of language and its enduring impact on our lives. In this appraisal, we shall explore the book's central themes, evaluate its distinctive writing style, and gauge its pervasive influence on the hearts and minds of its readership.

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