

Gay-Lussac's Law

- The gas in an aerosol can is at a temperature of 27°C and a pressure of 5 atm. If the gas reaches a pressure of 9 atm, what will the can's temperature be?
- An oxygen tank reads 150 kPa with a temperature of 25°C . What will the pressure read if the tank is cooled to -183°C ?

Combined Gas Law

- A gas has a volume of 35 liters at STP. What will its volume be at 5.0 atm and 35°C ?
- 0.4 L of a gas is contained at 29 atm and 273 K. What will the volume of the gas be at 15 atm and 293K?
- 400 mL of gas is stored at STP. The volume changes to 560 mL, and the temperature to 30°C . What is the pressure (in atm)?

Ideal Gas Law

- If 6 moles of a gas is at a pressure of 6.6 atm and a volume of 12 liters, what is the temperature of the gas?
- If there is unknown quantity of gas at a pressure of 122.6 kPa, a volume of 23 liters, and a temperature of 87°C , how many moles of gas are in the sample?
- If there are 5 moles of gas in a 50-liter container, at a temperature of 250 K, what is the pressure of the gas within the container?
- If 10.5 moles of gas are stored at a pressure of 0.15 atm and a temperature of 25°C , what is the volume of the container that the gas is stored in?

Dalton's Law

- A storage container holds a mixture of oxygen, helium, and carbon dioxide. If the partial pressures of the three gases tank are 25 atm of O_2 , 8 atm of He, and 25 atm of CO_2 , what is the total pressure inside of the container?
- The total pressure of a gas mixture is 1.9 kPa. The partial pressure of gas A is 0.06 kPa and the partial pressure of gas B is 0.25 kPa. What is the partial pressure of the remaining gas in the mixture?
- If I place 20 moles of N_2 and 1.5 moles of CO_2 in a 40 L container at a temperature of 25°C , what will the pressure of the resulting mixture of gases be?

Gas Laws Test Answer Key

JS Bruner



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