

# Gas Laws Study Guide

Write the Law, Formula, Symbols, Units, and a sentence describing the meaning of the symbols and units.

1. A sample of gas is held at a constant volume and pressure.
2. The pressure of a gas is held constant and the volume is held constant.
3. A sample of gas is held at a constant volume and temperature.
4. A sample of gas is held at a constant pressure and temperature.
5. The volume of a gas is held constant and the pressure is held constant.
6. A sample of gas is held at a constant volume and the pressure is held constant.

Name \_\_\_\_\_ Date \_\_\_\_\_

## Gas Laws Study Guide

Boyle's Temperature:  $T = \dots$   $K$   $T = \dots$   $^{\circ}C$

Boyle's Pressure:  $P = \dots$  atm  $P = \dots$  mmHg  $P = \dots$  kPa  $P = \dots$  bar  $P = \dots$  Pa

Boyle's Volume:  $V = \dots$  L  $V = \dots$  mL

### Boyle's Law

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$$

Combined Gas Law

$$P_{total} = P_1 + P_2 + P_3$$

Charles' Law

$$P_1 V_1 = P_2 V_2$$

Gay-Lussac's Law

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$$

Boyle's Law

$$P_1 T_1 = P_2 T_2$$

Boyle's Law

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$$

Boyle's Law

### Complete the following statements by filling in the blanks & writing the correct units.

As Pressure increases Volume decreases (increases) given that \_\_\_\_\_ is constant meaning that they have a direct (inverse) relationship.

As Volume increases Temperature increases (decreases) given that \_\_\_\_\_ is constant meaning that they have a direct (inverse) relationship.

As Pressure increases Temperature increases (decreases) given that \_\_\_\_\_ is constant meaning that they have a direct (inverse) relationship.

As the number of moles of gas in a container increases the Volume increases (decreases), the Pressure increases (decreases), and the Temperature increases (decreases).

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**Barron's Educational Series, Mark  
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