

Chapter 3 States of Matter

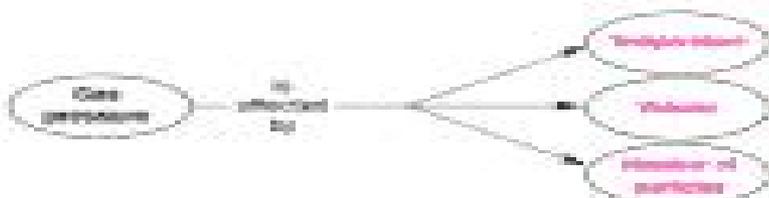
Section 3.2 The Gas Laws

(pages 75–81)

This section discusses gas pressure and the factors that affect it. It also explains the relationships between the temperature, volume, and pressure of a gas.

Reading Strategy (page 75)

Identifying Cause and Effect As you read, identify the variables that affect gas pressure, and write them in the diagram below. For more information on this Reading Strategy, see the **Reading and Study Skills** in the **Skills and Reference Handbook** at the end of your textbook.



Pressure (pages 75–76)

1. What is pressure? Pressure is a result of a force distributed over an area.
2. Circle the letter of each unit used to express amounts of pressure.

a. newton	b. joule
<input checked="" type="checkbox"/> Pascal	<input checked="" type="checkbox"/> kilopascal
3. What causes the pressure in a closed container of gas? Collisions between particles of gas and the walls of the container cause the pressure.

Factors that Affect Gas Pressure (pages 76–77)

4. Name the factors that affect the pressure of an enclosed gas.

a. <u>the temperature</u>	b. <u>the volume</u>	c. <u>The number of its particles</u>
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5. Is the following sentence true or false? In a closed container, increasing the temperature of a gas will decrease the force with which particles hit the walls of the container. False
6. What effect does raising the temperature of a gas have on its pressure, if the volume of the gas and the number of its particles are kept constant? The pressure of the gas will increase.
7. How does reducing the volume of a gas affect its pressure if the temperature of the gas and the number of particles are constant? Reducing the volume of a gas increases its pressure if the temperature and volume are constant.
8. Increasing the number of particles of a gas will increase its pressure if the temperature and the volume are constant.

Gas Laws Review Answer Key

K Morrison



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