

**12 • The Gas Laws****THE COMBINED GAS LAW**

In practical terms, it is often difficult to hold any of the variables constant. When there is a change in pressure, volume and temperature, the combined gas law is used.

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2} \quad \text{or} \quad P_1 \times V_1 \times T_2 = P_2 \times V_2 \times T_1$$

$$K = ^\circ C + 273$$

*See  
scratch  
paper*

Complete the following chart.

	$P_1$	$V_1$	$T_1$	$P_2$	$V_2$	$T_2$
1	1.50 atm	3.00 L	<del>20.0°C</del> 293 K	2.50 atm		<del>30.0°C</del> 303 K
2	720. torr	256. mL	<del>25.0°C</del> 298 K		250. mL	<del>50.0°C</del> 323 K
3	600. mmHg	2.50 L	<del>22.0°C</del> 325 K	760. mmHg	1.80 L	
4		750. mL	<del>0.00°C</del> 273 K	2.00 atm	500. mL	<del>25.0°C</del> 298 K
5	95.0 kPa	4.00 L		101. kPa	6.00 L	<del>498.°C</del> 471. K or
6	650. torr		<del>100.°C</del> 373 K	900. torr	225. mL	<del>150.°C</del> 423 K
7	850. mmHg	1.50 L	<del>15.0°C</del> 288 K		2.50 L	<del>30.0°C</del> 303 K
8	125. kPa	125. mL		100. kPa	100 mL	<del>75.0°C</del> 348 K

# Gas Laws Practice Packet

**S Baum**



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