

Gas Variables

How are the variables that describe a gas related?

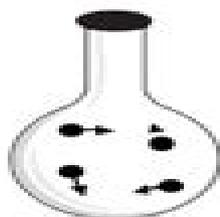
Why?

Imagine buying a balloon bouquet at a party store. How will the helium gas in the bouquet behave if you carry it outside on a hot summer day? How will it behave if you carry it outside during a snowstorm?

What happens if the balloons are made of latex, which can stretch? What happens if the balloons are made of Mylar[®], which cannot stretch? What if you add just a small amount of gas to each balloon? What if you add a lot of gas? In this activity, you will explore four variables that quantify gases—pressure (P), volume (V), temperature (T), and moles (n) of gas. These four variables can be related mathematically so that predictions about gas behavior can be made.

Model 1 - Gases in a Nonflexible Container

Experiment A (Adding more gas)



A1

Volume = 1 unit
External pressure = 1 atm
Internal pressure = 1 atm
Temperature = 200 K



A2

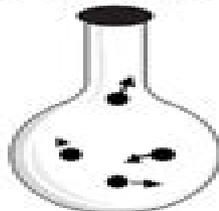
Volume = 1 unit
External pressure = 1 atm
Internal pressure = 2 atm
Temperature = 200 K



A3

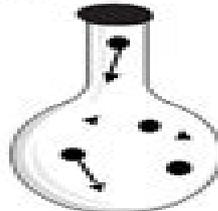
Volume = 1 unit
External pressure = 1 atm
Internal pressure = 3 atm
Temperature = 200 K

Experiment B (Heating the gas)



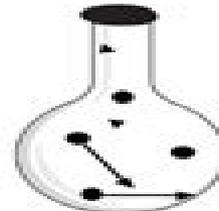
B1

Volume = 1 unit
External pressure = 1 atm
Internal pressure = 1 atm



B2

Volume = 1 unit
External pressure = 1 atm
Internal pressure = 2 atm



B3

Volume = 1 unit
External pressure = 1 atm
Internal pressure = 3 atm

Gas Variables Packet

Rachel S Tattersall



Gas Variables Packet:

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