

Unit 7 – Equilibrium (7-9%)

Some reactions can occur in both forward and reverse directions, sometimes proceeding in each direction simultaneously.

7.1 Intro to Equilibrium: Explain the relationship between the occurrence of a reversible chemical or physical process, and the establishment of equilibrium, to experimental observations.

- Many observable processes are reversible. Examples include evaporation and condensation of water, absorption and desorption of a gas, or dissolution and precipitation of a salt. Some important reversible chemical processes include the transfer of protons in acid-base reactions and the transfer of electrons in redox reactions.
- When equilibrium is reached, no observable changes occur in the system. Reactants and products are simultaneously present, and the concentrations or partial pressures of all species remain constant.
- The equilibrium state is dynamic. The forward and reverse processes continue to occur at equal rates, resulting in no net observable change.
- Graphs of concentration, partial pressure, or rate of reaction versus time for simple chemical reactions can be used to understand the establishment of chemical equilibrium.

7.2 Direction of Reversible Reactions: Explain the relationship between the direction in which a reversible reaction proceeds and the relative rates of the forward and reverse reactions.

Unit 7 Objectives

Equilibrium

A system at equilibrium depends on the relationships between concentrations, partial pressures of chemical species, and equilibrium constant K .

AP Chemistry

REFERENCE



Modeling Chemistry Unit 7 Objectives

ML Morrison



Modeling Chemistry Unit 7 Objectives:

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