



# COVALENT BONDING

## Practice Problems

Do your answers to the following problems.

### SECTION 8.1 MOLECULAR COMPOUNDS

- Classify each of the following as either ionic or molecular.
  - $\text{NaCl}$
  - $\text{H}_2\text{O}$
  - $\text{HCl}$
  - $\text{KNO}_3$
  - $\text{H}_2\text{SO}_4$
  - $\text{H}_2\text{O}_2$
  - $\text{C}_2\text{H}_6$
  - $\text{C}_2\text{H}_4$
  - $\text{C}_2\text{H}_2$

- Classify each of the following as either ionic or molecular compound. Write the name of the compound.
- $\text{K}_2\text{SO}_4$
- $\text{C}_2\text{H}_6$
- $\text{C}_2\text{H}_4$
- $\text{C}_2\text{H}_2$
- $\text{C}_2\text{H}_6\text{O}$
- $\text{C}_2\text{H}_4\text{O}$
- $\text{C}_2\text{H}_2\text{O}$

### SECTION 8.2 THE NATURE OF COVALENT BONDING

- Show covalent bonding in hydrogen fluoride.  $\text{H}-\text{F}$
- Show covalent bonding in methane.  $\text{CH}_4$
- Show covalent bonding in carbon dioxide.  $\text{CO}_2$
- Show covalent bonding in acetylene.  $\text{H}-\text{C}\equiv\text{C}-\text{H}$
- Draw Lewis structures for the following molecules.
  - $\text{C}_2\text{H}_6$
  - $\text{C}_2\text{H}_4$
  - $\text{C}_2\text{H}_2$
  - $\text{C}_2\text{H}_6\text{O}$
  - $\text{C}_2\text{H}_4\text{O}$
  - $\text{C}_2\text{H}_2\text{O}$



### SECTION 8.3 BONDING: THE ORBITALS

- Which orbital and sub-orbital are represented by each of the following?
  - $2s$
  - $2p$
  - $3s$
  - $3p$
  - $3d$
- Which orbital and sub-orbital give rise to each of the following?
  - $2s$
  - $2p$
  - $3s$
  - $3p$
  - $3d$

# Guided Practice Problem Chapter 8 Covalent Bonding

**Anthony S. Fauci**



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## **Table of Contents Guided Practice Problem Chapter 8 Covalent Bonding**

1. Understanding the eBook Guided Practice Problem Chapter 8 Covalent Bonding
  - The Rise of Digital Reading Guided Practice Problem Chapter 8 Covalent Bonding
  - Advantages of eBooks Over Traditional Books
2. Identifying Guided Practice Problem Chapter 8 Covalent Bonding
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Guided Practice Problem Chapter 8 Covalent Bonding
  - User-Friendly Interface
4. Exploring eBook Recommendations from Guided Practice Problem Chapter 8 Covalent Bonding
  - Personalized Recommendations
  - Guided Practice Problem Chapter 8 Covalent Bonding User Reviews and Ratings
  - Guided Practice Problem Chapter 8 Covalent Bonding and Bestseller Lists
5. Accessing Guided Practice Problem Chapter 8 Covalent Bonding Free and Paid eBooks
  - Guided Practice Problem Chapter 8 Covalent Bonding Public Domain eBooks
  - Guided Practice Problem Chapter 8 Covalent Bonding eBook Subscription Services
  - Guided Practice Problem Chapter 8 Covalent Bonding Budget-Friendly Options

6. Navigating Guided Practice Problem Chapter 8 Covalent Bonding eBook Formats
  - ePub, PDF, MOBI, and More
  - Guided Practice Problem Chapter 8 Covalent Bonding Compatibility with Devices
  - Guided Practice Problem Chapter 8 Covalent Bonding Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Guided Practice Problem Chapter 8 Covalent Bonding
  - Highlighting and Note-Taking Guided Practice Problem Chapter 8 Covalent Bonding
  - Interactive Elements Guided Practice Problem Chapter 8 Covalent Bonding
8. Staying Engaged with Guided Practice Problem Chapter 8 Covalent Bonding
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Guided Practice Problem Chapter 8 Covalent Bonding
9. Balancing eBooks and Physical Books Guided Practice Problem Chapter 8 Covalent Bonding
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Guided Practice Problem Chapter 8 Covalent Bonding
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Guided Practice Problem Chapter 8 Covalent Bonding
  - Setting Reading Goals Guided Practice Problem Chapter 8 Covalent Bonding
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Guided Practice Problem Chapter 8 Covalent Bonding
  - Fact-Checking eBook Content of Guided Practice Problem Chapter 8 Covalent Bonding
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

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