

Half-Lives

1. Radioactive materials undergo nuclear decay at a constant rate.
2. The half-life is the time required for half of a sample to undergo radioactive decay.
3. The half-life of iodine - 131 is about 8 days. This means in eight days, half of a sample would remain.

a. Start with 200 mg of iodine - 131. How many remains after eight days? 100 mg

$$200 \text{ mg} \xrightarrow[HL]{8d} 100 \text{ mg}$$

b. You have 50 mg of iodine - 131 remaining. How many half-lives have passed? 2

How many days is this? 16

$$100\% \rightarrow 50\% \rightarrow 25\% \rightarrow 12.5\%$$

c. How long would it take for 87.5% of iodine to decay?

$$3(\text{HL}) = 24 \text{ days}$$

Half-Life Formula:

$$A = A_0 \left(\frac{1}{2}\right)^n$$

$$n = t/h$$

A = amount left

t = total time passing

A₀ = initial amount

h = time for one half-life

n = number of half-lives

1. Label the variables above.
2. Nobelium - 259 has a half-life of 58 minutes. How much remains from a 1000g sample after 1440 minutes?

$$n = \frac{t}{h} = \frac{1440}{58} = 24.8$$

$$A = 1000 \left(\frac{1}{2}\right)^{24.8}$$

$$\rightarrow 3.36 \times 10^{-5} \text{ g remaining}$$

3. Dubnium - 269 has a half-life of 34 seconds. How long will you wait to see 500g of it decay to 1g?

$$A = A_0 \left(\frac{1}{2}\right)^n$$

$$\log(0.002) = \log(0.5)^{\frac{t}{34}}$$

$$t = 305 \text{ s}$$

$$\frac{1}{500} = \frac{500 \left(\frac{1}{2}\right)^{\frac{t}{34}}}{500} \rightarrow 0.002 \left(\frac{1}{2}\right)^{\frac{t}{34}} \rightarrow \frac{-2.7}{-0.3} = \frac{-0.3 \left(\frac{t}{34}\right)}{-0.3}$$

$$(34) 8.97 = \left(\frac{t}{34}\right)^{34}$$

Practice Problems using the Formula (SHOW YOUR WORK)

1. The half-life of polonium - 218 is 3 min. If you start with 485g of polonium, how much remains after 22 minutes?

$$A = 485 \left(\frac{1}{2}\right)^{\frac{22}{3}} = 485 \left(\frac{1}{2}\right)^{7.3} = 3.08 \text{ g left}$$

2. The half-life of polonium - 218 is 3 min. How much time will it take to see 25g decay to 3g?

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Table of Contents Guided Practice Problems Answers

1. Understanding the eBook Guided Practice Problems Answers
 - The Rise of Digital Reading Guided Practice Problems Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Guided Practice Problems Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Guided Practice Problems Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Guided Practice Problems Answers
 - Personalized Recommendations
 - Guided Practice Problems Answers User Reviews and Ratings
 - Guided Practice Problems Answers and Bestseller Lists
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 - Adjustable Fonts and Text Sizes of Guided Practice Problems Answers
 - Highlighting and Note-Taking Guided Practice Problems Answers
 - Interactive Elements Guided Practice Problems Answers
- 8. Staying Engaged with Guided Practice Problems Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Guided Practice Problems Answers
- 9. Balancing eBooks and Physical Books Guided Practice Problems Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Guided Practice Problems Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Guided Practice Problems Answers
 - Setting Reading Goals Guided Practice Problems Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Guided Practice Problems Answers
 - Fact-Checking eBook Content of Guided Practice Problems Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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